

3. a) The mean kinetic energy of a nitrogen molecule is the same as the mean kinetic energy of an oxygen molecule since E_k is directly proportional to temperature $E_k = \frac{3}{2} kT$

b) molar mass $N_2 = 0.028 \text{ kg mol}^{-1}$ molar mass $O_2 = 0.032 \text{ kg mol}^{-1}$
 $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$

$$\frac{1}{2} M_N C_{\text{rms},N}^2 = \frac{1}{2} M_O C_{\text{rms},O}^2$$

$$\frac{0.028}{6.02 \times 10^{23}} C_{\text{rms},N}^2 = \frac{0.032}{6.02 \times 10^{23}} C_{\text{rms},O}^2$$

$$0.028 C_{\text{rms},N}^2 = 0.032 C_{\text{rms},O}^2$$

$$C_{\text{rms},N}^2 = \frac{0.032}{0.028} C_{\text{rms},O}^2$$

$$C_{\text{rms},N}^2 = 1.14 C_{\text{rms},O}^2$$

$$C_{\text{rms},N} = \sqrt{1.14} C_{\text{rms},O}$$

$$C_{\text{rms},N} = 1.07 C_{\text{rms},O}$$

4. molar mass = $0.028 \text{ kg mol}^{-1}$ $V = 0.037 \text{ m}^3$ $T = 300 \text{ K}$
 $p = 100 \text{ kPa} = 100 \times 10^3 \text{ Pa}$

a) $pV = nRT \Rightarrow \frac{pV}{RT} = n = \frac{(100 \times 10^3)(0.037)}{8.31(300)}$
 $= \frac{3700}{2493} = 1.484 = \underline{1.48 \text{ moles to 3sf}}$

b) mass = number of moles \times molar mass
 $= 1.484 \times 0.028 = 0.041512 \text{ kg}$
 $= \underline{4.2 \times 10^{-2} \text{ kg to 2sf}}$

(If 1.48 moles used then mass = $0.04144 \text{ kg} = 4.1 \times 10^{-2} \text{ kg to 2sf}$)

c) mass of one molecule $m = \frac{\text{molar mass}}{N_A} = \frac{0.028}{6.02 \times 10^{23}} = 4.65 \times 10^{-26} \text{ kg}$

$$\frac{1}{2} M C_{\text{rms}}^2 = \frac{3}{2} kT$$

$$4.65 \times 10^{-26} C_{\text{rms}}^2 = 3 \left(1.38 \times 10^{-23} \right) 300 = 1.242 \times 10^{-20}$$

$$C_{\text{rms}} = \frac{1.242 \times 10^{-20}}{4.65 \times 10^{-26}} = 267 \times 10^3 \text{ m s}^{-1}$$

$$C_{\text{rms}} = \sqrt{267 \times 10^3}$$

$$= 516 \text{ m s}^{-1}$$

$$= \underline{5.2 \times 10^2 \text{ m s}^{-1} \text{ to 2sf}}$$

w.3 could use
 $pV = \frac{1}{2} N m C_{\text{rms}}^2$
 Since $Nm = \text{mass}$
 and would also be
 equal to 1.48(0.028)

Conceptual Physics Chapter 20 Gases Exercises Answers

Raymond A. Serway



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Physics, Volume 1 Robert Resnick, David Halliday, Kenneth S. Krane, 2002 Each chapter includes questions and problems

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Table of Contents Conceptual Physics Chapter 20 Gases Exercises Answers

1. Understanding the eBook Conceptual Physics Chapter 20 Gases Exercises Answers
 - The Rise of Digital Reading Conceptual Physics Chapter 20 Gases Exercises Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Conceptual Physics Chapter 20 Gases Exercises Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Conceptual Physics Chapter 20 Gases Exercises Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conceptual Physics Chapter 20 Gases Exercises Answers
 - Personalized Recommendations
 - Conceptual Physics Chapter 20 Gases Exercises Answers User Reviews and Ratings
 - Conceptual Physics Chapter 20 Gases Exercises Answers and Bestseller Lists

5. Accessing Conceptual Physics Chapter 20 Gases Exercises Answers Free and Paid eBooks
 - Conceptual Physics Chapter 20 Gases Exercises Answers Public Domain eBooks
 - Conceptual Physics Chapter 20 Gases Exercises Answers eBook Subscription Services
 - Conceptual Physics Chapter 20 Gases Exercises Answers Budget-Friendly Options
6. Navigating Conceptual Physics Chapter 20 Gases Exercises Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Conceptual Physics Chapter 20 Gases Exercises Answers Compatibility with Devices
 - Conceptual Physics Chapter 20 Gases Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Highlighting and Note-Taking Conceptual Physics Chapter 20 Gases Exercises Answers
 - Interactive Elements Conceptual Physics Chapter 20 Gases Exercises Answers
8. Staying Engaged with Conceptual Physics Chapter 20 Gases Exercises Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Conceptual Physics Chapter 20 Gases Exercises Answers
9. Balancing eBooks and Physical Books Conceptual Physics Chapter 20 Gases Exercises Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conceptual Physics Chapter 20 Gases Exercises Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Conceptual Physics Chapter 20 Gases Exercises Answers
 - Setting Reading Goals Conceptual Physics Chapter 20 Gases Exercises Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Fact-Checking eBook Content of Conceptual Physics Chapter 20 Gases Exercises Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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