

- **Matter** = anything that has mass and volume (takes up space)
- **Mass** = the quantity of matter in an object (also, how much force it takes to accelerate it)
- **Volume** = the amount of space occupied by an object

I Properties of Matter

- A. **Physical Properties** – identifying characteristics that can be determined without changing the composition of the material. Ex. – boiling point, volume, solubility, color, Sample (how much you have). Ex. – length, quantity of heat, mass, volume.
- B. **Chemical Properties** – identifying characteristics of a material that can only be determined by an attempt to change the composition of the material; describes the ability to be changed, not the actual process of changing. Ex. – flammability, oxidation, reaction with acids.
- C. **States of Matter** – physical form of materials.
 1. **Solids** – state of matter that has both a definite volume and a definite shape.
 2. **Liquids** – state of matter that can flow and has a definite volume but assumes the shape of its container.
 3. **Gases** – state of matter that can flow and has neither a definite volume nor a definite shape.

II Changes in Matter

- A. **Physical Changes** – alteration to materials in which shape, number of pieces, or state is different but the identifying characteristics remain the same. Ex. – melting, bending, breaking, dissolving.
- B. **Chemical Changes** – alteration to materials in which the composition of the material is different; a new material is produced; describes the actual process of change. Ex. – burning, decomposing, explosions, rusting, any reaction which makes a new material.
- C. **Signs of a Chemical Change** (What to Look for in Labs):
 - a. Change in color
 - b. Change in odor
 - c. Production of a new substance
 - i. Gas (seen as bubbles)
 - ii. Solid (seen as a "precipitate" or other new, solid substance)
 - d. Energy change that did not result from adding or removing heat/energy manually
 - i. An increase in temperature indicates that heat was released by the reaction to the environment (exothermic)
 - ii. A decrease in temperature indicates that heat was absorbed by the reaction from the environment (endothermic)

IV Classifying Matter

- **Pure substance** = material that cannot be physically separated into its components. Elements and compounds (molecules and ionic compounds) are pure substances.
 - **Mixture** = material that can be physically separated into its components
- A. **Mixtures** – type of matter that contains two or more materials that are not chemically bonded and are not in a definite ratio (proportion).
1. **Heterogeneous Mixtures** – Composition is NOT the same throughout. Matter that contains two or more materials (not chemically bonded) and has visually distinguishable parts. Ex. – soil, vegetable soup, beach sand.
 2. **Homogeneous Mixtures (Solutions)** – Composition is the same throughout. Matter that contains two or more materials (not chemically bonded) and has no visually distinguishable parts. Ex. – soda, air, brass.

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