

Chemistry: Periodic Table Unit Study Guide

Objectives:

1. Define the following terms:
 - Period
 - Group
 - Atomic Number
 - Mass Number
2. State the Periodic Law.
3. Explain the following trends:
 - Atomic Radius
 - Electronegativity
 - Ionization Energy
 - Electron Affinity
4. An element has the following properties:
 - Atomic Number: 19
 - Mass Number: 39
 - Group: 1
 - Period: 4
5. Define the following terms:
 - Alkali Metals
 - Alkaline Earth Metals
 - Transition Metals
 - Halogens
 - Noble Gases
6. Define the following terms:
 - Metals
 - Nonmetals
 - Metalloids
7. Classify the following elements as metals, nonmetals, or metalloids:
 - Sodium (Na)
 - Chlorine (Cl)
 - Iron (Fe)
 - Carbon (C)
 - Aluminum (Al)
 - Sulfur (S)
 - Gold (Au)
 - Hydrogen (H)
 - Helium (He)
 - Calcium (Ca)
 - Oxygen (O)
 - Potassium (K)
 - Nitrogen (N)
 - Copper (Cu)
 - Phosphorus (P)
 - Silver (Ag)
 - Argon (Ar)
 - Barium (Ba)
 - Fluorine (F)
 - Zinc (Zn)
 - Neon (Ne)
 - Lead (Pb)
 - Bromine (Br)
 - Mercury (Hg)
 - Krypton (Kr)
 - Strontium (Sr)
 - Iodine (I)
 - Thallium (Tl)
 - Xenon (Xe)
 - Bismuth (Bi)
 - Polonium (Po)
 - Astatine (At)
 - Francium (Fr)
 - Radium (Ra)
 - Actinium (Ac)
 - Protactinium (Pa)
 - Uranium (U)
 - Neptunium (Np)
 - Plutonium (Pu)
 - Americium (Am)
 - Curium (Cm)
 - Berkelium (Bk)
 - Californium (Cf)
 - Einsteinium (Es)
 - Fermium (Fm)
 - Mendelevium (Md)
 - Nobelium (No)
 - Lanthanum (La)
 - Cerium (Ce)
 - Praseodymium (Pr)
 - Neodymium (Nd)
 - Europium (Eu)
 - Gadolinium (Gd)
 - Terbium (Tb)
 - Dysprosium (Dy)
 - Homium (Ho)
 - Erbium (Er)
 - Thulium (Tm)
 - Ytterbium (Yb)
 - Lutetium (Lu)
8. Sketch a periodic table showing the location of each of the following groups:
 - Alkali Metals
 - Alkaline Earth Metals
 - Transition Metals
 - Halogens
 - Noble Gases
9. Fill in the following chart:

Name: _____ Date: _____ Period: _____

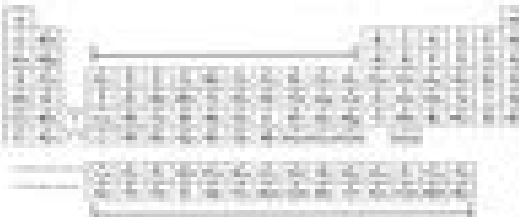
Periodic Table Unit Study Guide

Objectives and Overview of the Periodic Table

1. What are the periodic trends in the periodic table?

- Atomic radius increases from top-left to bottom-right.
- Electronegativity increases from bottom-left to top-right.
- Ionization energy increases from bottom-left to top-right.
- Electron affinity increases from bottom-left to top-right.

2. Label the following groups on the periodic table: Alkali Metals, Alkaline Earth Metals, Transition Metals, Halogens, and Noble Gases. For each group, write down the color designation and the number of valence electrons.



Alkali Metals

Alkaline Earth Metals


Transition Metals

Halogens

Noble Gases

3. Sketch the periodic table showing the location of each of the following groups: Alkali Metals, Alkaline Earth Metals, Transition Metals, Halogens, and Noble Gases.

4. Fill in the following chart:

Element	Astatine	Ca ²⁺	1s ² 2s ² 2p ⁶ 3s ²	
Number of Protons				
Number of Electrons				
Atomic Number (Z)				
Mass Number (A)				

Chemistry The Periodic Table Study Guide

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