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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ 2 $\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + 3\text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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$$\text{br } 1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3$$

$$\text{br } 2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o}$$

$$\text{br } 3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$$

$$\text{br } 4 \text{ 2 ga} + \text{br}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2 + \text{so}_3 + 3 \text{ 6 nabr}$$

$$\text{br } 5 \text{ 3 sno} + 2 \text{ nf}_3 + 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3$$
 using the following equation

$$\text{br } 2 \text{ naoh} + \text{h}_2\text{so}_4 + 2 \text{ h}_2\text{o} + \text{na}_2\text{so}_4$$

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$$1 \text{ 2 3 4 5 i n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3$$

$$2 \text{ cah} + 40 \text{ 17 o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} + \text{hbr} + \text{khco}_3 + \text{h}_2\text{o} + \text{kbr} + \text{co}_2 + \text{gabr} + \text{na}_2\text{so}_3 + \text{gaz} + \text{so}_3 + 3 \text{ nabr} + \text{sno} + \text{nf}_3 + \text{snf}_2$$

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 answer: $\text{ce SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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$$\text{n}_2 + \text{f}_2 + \text{nf} + \text{c}_6\text{h}_{10} + 02 \text{ c}_0 + \text{ho} + 3 \text{ hbr} + \text{khco}_3 + \text{h}_2\text{o} + \text{kbr} + 4 \text{ gabry} + \text{na} + \text{so}_3 + \text{gaz} + \text{so}_3 + 3 \text{ 5 sno} + \text{nf} + \text{snf}_2 + \text{n}_2\text{o} + \text{co}_2 + \text{nabr}$$
 solve the following stoichiometry grams grams problems 6

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