

Concentrations and Dilutions: How to Dilute a Concentrated Solution

In the first part of this lab, you learned about the actions of solutes added to water to form solutions of various concentrations. In this exercise, you will use a powerful and simple tool to determine the concentration of a solution when it has been diluted with additional water or made more concentrated due to evaporation of water.

Some basic equations for dilution effects:

Molarity:

Dilution:

Evaporation:

Procedure: Print this page with the other MP Chemistry MP Concentrations [Worksheet](#)

Point 1: Dilution

At this point you should be familiar with the Concentration simulation. Because some items may have different labels or have mistakes in the original simulation, pay particular attention to the effect of evaporation and addition of water to the solution's concentration.



How does the concentration change as additional water is added?

Why?



How does the concentration change as evaporation occurs?

Why?

Why is this?

Does evaporation change the concentration of a concentrated solution?

Why is this?



Using the concentrated solution again, add 10 L of DI water to an empty beaker. What is the concentration?

Is this solution saturated?

How do you know?

Fill the beaker with another 10 L of water. What is the new concentration?

Complete the table below, using **Concentrations** as an strategy device, writing the new solution in the boxes provided.

Dilute 0.25 L of 25% solution	25% solution + 25 L water	0.25 L original + 25 L water	25% solution = 25 L solution

Repeat the exercise, using **Postulates** instead of an empty beaker.

Dilute 0.25 L of 25% solution	25% solution + 25 L water	0.25 L original + 25 L water	25% solution = 25 L solution

What do you notice about the concentration change as each addition of 0.25 L of water is added to the concentrated 25% solution?



Concentrated water

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