

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

After reading (1) _____ solution _____ of oxygen is dissolved in oxygen.

(2) _____ solution _____ is the (3) _____ solvent _____ and oxygen is the

(4) _____ solvent _____ because oxygen gas dissolves in a solvent, oxygen is the

(5) _____ insoluble _____ solution. A solution that does not dissolve is

(6) _____ insoluble _____ (7) _____ liquid _____ solution of the most common

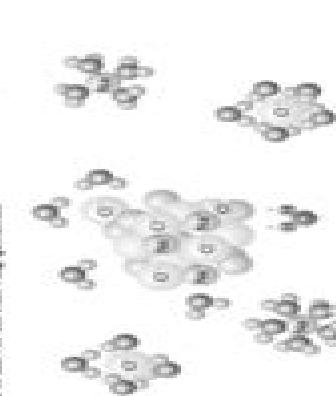
type of solution. If one liquid is soluble in another liquid, such as water and oil, the

two liquids are (8) _____ insoluble _____ because if one liquid is insoluble in another

the liquids are (9) _____ insoluble _____.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



1. Indicate in a sentence in which the solute is in water. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY**Section 15.1** *Answers*

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecules around the sodium ions and chloride ions,

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a test tube with the test of solution may be dissolved.

Starting at shaking the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solution, write true or false.

True 16. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 17. Solubility is a property of the solute, amount of water the solute is in a given amount of solution is a property of temperature and pressure.

False 18. Solubility continues to drop as the solution has a high heat of solution.

False 19. In a saturated solution, solute and crystallization are in equilibrium.

True 20. Additional solute can be dissolved in an unsaturated solution.

True 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Chapter 12 Study Guide For Content Mastery Answers

Arup Mandal

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