

CHAPTER 15 **STUDY GUIDE****Energy and Chemical Change****Section 15.1 Energy***In your textbook, read about the nature of energy.***In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.**

1. Energy is the ability to do work or produce heat.
2. The law of conservation of energy states that energy *can be* created and destroyed.
3. Chemical potential energy is energy stored in a substance because of its *composition*.
4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
6. A *calorie* is the SI unit of heat and energy.
7. The *specific* heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
8. *Kinetic* energy is energy of motion.
9. Chemicals participating in a chemical reaction contain only *potential energy*.
10. One nutritional Calorie is equal to *100 calories*.
11. One calorie equals *4.184 joules*.
12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Chapter 15 Study Guide Answer Key

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Rodgers, David Wilson**

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