

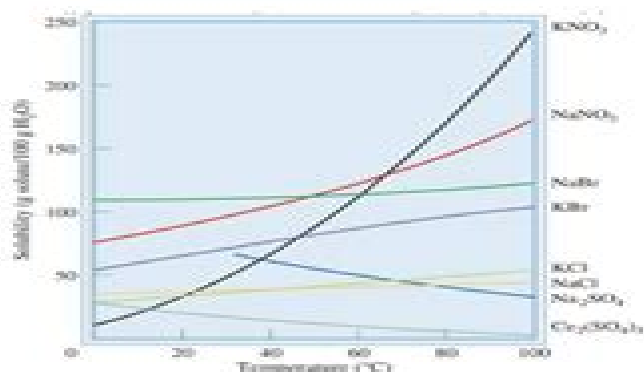
General Chemistry II Chapter 13 Study Guide

Concentration units review:

- Dilute solutions: contain small amounts of solute
- Concentrated solutions: contain large amounts of solute
- Molarity changes with temperature; molality is temperature-independent

Solubility of solids as a function of temperature:

- For most solids, as temperature increases, solubility increases



Solubility of gasses: effect of temperature:

- Gasses are less soluble in liquids as temperature increases
- Under high temperatures:
 - Molecules move faster, disrupt intermolecular forces, gasses escape the liquid and fewer gas molecules/atoms in liquid = lower solubility
- In a sealed container, pressure will increase with increasing temperature; more gas out of the solution, moving faster, colliding more with inner walls of container

Solubility of gasses: effect of pressure:

- Gas solubility in liquids increases as pressure increases; gas molecules pushed are too close together so they are forced back into the liquid
- Henry's law: $c = kP$
 - c = conc. (M) of the dissolved gas
 - P = external pressure (atm)
 - k = Henry's law constant (unit: mol/L x atm); is different for each solute-solvent pair, varies with temperature
 - If you know the conc. At one pressure, determined k allows you to determine conc. At different pressures

Chapter 13 Study Guide Chemistry Answers

Martin Silberberg, Libby Weberg



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Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove, Robert L. Caret, 1981

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