

20.3 The kinetic theory of gases Page 126

- 1a) Pressure increases when temperature increases ..
- the speed of the particles increases because the mean kinetic energy of a molecule $\propto \frac{1}{2} kT$. So, as temperature increases the kinetic energy of the molecules increases and since $E_k = \frac{1}{2}mv^2$ the speed of the molecules increases
 - if the speed of the molecules increases then the collisions with the wall of the container will be more frequent and the impacts will be harder

b) O_2 molar mass $= 0.032 \text{ kg mol}^{-1}$ $V = 0.025 \text{ m}^3$ $T = 373 \text{ K}$ $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$
 $p = 120 \text{ kPa} = 120 \times 10^3 \text{ Pa}$

i) $pV = nRT \Rightarrow \frac{pV}{RT} = n = \frac{120 \times 10^3 (0.025)}{8.31 (373)} = \frac{3000}{3099.43} = 0.96786$

ii) The total kinetic energy of a mole of an ideal gas $= \frac{3}{2} nRT = \frac{3}{2} (0.97) (8.31) (373)$
 $= 4507.96 \text{ J} = 4.5 \times 10^3 \text{ J}$ to 2sf

OR

The total kinetic energy of N particles $= N \cdot \frac{3}{2} kT = N \cdot \frac{3}{2} \frac{R}{N_A} T = \frac{3}{2} nRT$

2. H_2 molar mass $= 0.002 \text{ kg mol}^{-1}$ $T = 0^\circ \text{C} = 273 \text{ K}$ $N_A = 6.02 \times 10^{23}$

i) mean kinetic energy of a molecule $= \frac{3}{2} kT = \frac{3}{2} (1.38 \times 10^{-23}) (273)$
 $= 5.654 \times 10^{-21} \text{ J} = 5.7 \times 10^{-21} \text{ J}$ to 2sf

b) kinetic energy $= \frac{1}{2} m c_{rms}^2 = \frac{3}{2} kT = 5.654 \times 10^{-21} \text{ J}$

$\frac{1}{2} \frac{\text{molar mass}}{N_A} c_{rms}^2 = 5.654 \times 10^{-21} \text{ J}$

$\frac{1}{2} \left(\frac{0.002}{6.02 \times 10^{23}} \right) c_{rms}^2 = 5.654 \times 10^{-21} \text{ J}$

$\frac{1}{2} (3.32 \times 10^{-27}) c_{rms}^2 = 5.654 \times 10^{-21} \text{ J}$

$c_{rms}^2 = \frac{2 (5.654 \times 10^{-21})}{3.32 \times 10^{-27}} = 3.4041 \times 10^6$

$c_{rms} = \sqrt{3.4041 \times 10^6} = 1.845 \times 10^3 \text{ ms}^{-1} = 1.8 \times 10^3 \text{ ms}^{-1}$ to 2sf

$\frac{1}{2} 5.7 \times 10^{-21} \text{ J}$
 n and
 $c_{rms} = 1.85 \times 10^3 \text{ ms}^{-1}$
 $= 1.8 \times 10^3 \text{ ms}^{-1}$
 to 2sf

Chapter 20 Gases Exercises Answers

Louis J. Thibodeaux, Donald Mackay



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