

CHAPTER 9 REVIEW

Stoichiometry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ The coefficients in a chemical equation represent the
 - (a) masses in grams of all reactants and products.
 - (b) relative number of moles of reactants and products.
 - (c) number of atoms of each element in each compound in a reaction.
 - (d) number of valence electrons involved in a reaction.
2. _____ Which of the following would not be studied within the topic of stoichiometry?
 - (a) the mole ratio of Al to Cl in the compound aluminum chloride
 - (b) the mass of carbon produced when a known mass of sucrose decomposes
 - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
 - (d) the amount of energy required to break the ionic bonds in CaF_2
3. _____ A balanced chemical equation allows you to determine the
 - (a) mole ratio of any two substances in the reaction.
 - (b) energy released in the reaction.
 - (c) electron configuration of all elements in the reaction.
 - (d) reaction mechanism involved in the reaction.
4. _____ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - (a) reaction sequence.
 - (b) bond energy.
 - (c) mole ratio.
 - (d) element proportion.
5. Given the reaction represented by the following unbalanced equation:
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 - a. Balance the equation.

b. What is the mole ratio of NO_2 to O_2 ?

c. If 20.0 mol of NO_2 form, how many moles of O_2 must have been consumed?

d. Twice as many moles of NO_2 form as moles of N_2O are consumed. True or False?

e. Twice as many grams of NO_2 form as grams of N_2O are consumed. True or False?

Chapter 9 Reading Guide Stoichiometry

T. R. Dickson



Chapter 9 Reading Guide Stoichiometry:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students CliffsAP Chemistry, 4th Edition Bobrow Test Preparation Services, 2011-09-26 Your complete guide to a higher score on the AP Chemistry exam Why CliffsAP Guides Go with the name you know and trust Get the information you need fast Written by test prep specialists Contents include Introduction overview of the test and how it is scored proven strategies for each type of question Review of topics tested atom periodic table bonding geometry hybridization stoichiometry gases liquids and solids thermodynamics solutions equilibrium acids and bases kinetics redox nuclear chemistry organic chemistry and writing reactions The Labs feature 20 multiple choice questions multiple free response questions on each topic with answers on each topic with answers and explanations scoring rubrics and 2 full length practice exams Structured like the actual exam Complete with answers and explanations AP is a registered trademark of the College Board which was not involved in the production of and does not endorse this product **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 **Barron's Chemistry Practice Plus: 400+ Online Questions and Quick Study Review** Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2022-07-05 Barron's Chemistry Practice Plus features more than 400

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- Chapter 2: Essential Elements of Chapter 9 Reading Guide Stoichiometry
- Chapter 3: Chapter 9 Reading Guide Stoichiometry in Everyday Life
- Chapter 4: Chapter 9 Reading Guide Stoichiometry in Specific Contexts
- Chapter 5: Conclusion

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