

4. A sample of oxygen gas has a volume of 150 mL when its pressure is 440 mmHg. If the pressure is increased to standard pressure and the temperature remains constant, what will the new gas volume be?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$P_1 = 440 \text{ mmHg} \quad V_1 = 150 \text{ mL} \quad P_2 = 760 \text{ mmHg (1 atm)} \quad V_2 = ?$$

$$(440 \text{ mmHg})(150 \text{ mL}) = V_2 (760 \text{ mmHg})$$

$$V_2 = 87 \text{ mL}$$

5. Ralph had a helium balloon with a volume of 4.88 liters at 150 kPa of pressure. If the volume is changed to 3.15 liters, what would be the new pressure in atm?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$P_1 = 150 \text{ kPa} \quad P_2 = ?$$

$$V_1 = 4.88 \text{ L} \quad V_2 = 3.15 \text{ L}$$

$$(4.88 \text{ L})(150 \text{ kPa}) = (3.15 \text{ L}) P_2 \quad P_2 = 232 \text{ kPa} \times \frac{1 \text{ atm}}{101.3 \text{ kPa}}$$

$$P_2 = 2.29 \text{ atm}$$

6. 5.36 liters of nitrogen gas are at -25°C and 733 mm Hg. What would be the volume at 128°C and 1.5 atm?

Combined gas law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$P_1 = 733 \text{ mmHg} \times \frac{1 \text{ atm}}{760 \text{ mmHg}} = 0.964 \text{ atm}$$

$$V_1 = 5.36 \text{ L} \quad T_1 = -25^\circ\text{C} + 273 = 248 \text{ K}$$

$$P_2 = 1.5 \text{ atm} \quad T_2 = 128^\circ\text{C} + 273 = 401 \text{ K}$$

$$\frac{(0.964 \text{ atm})(5.36 \text{ L})}{248 \text{ K}} = \frac{V_2 (1.5 \text{ atm})}{401 \text{ K}} \quad V_2 = 5.57 \text{ L}$$

7. At constant temperature, 245 mL of a gas at 4 atm of pressure is expanded to 1.75 L. What is the new pressure?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$V_1 = 245 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = \text{approx} 0.245 \text{ L}$$

$$P_1 = 4 \text{ atm}$$

$$V_2 = 1.75 \text{ L}$$

$$(0.245 \text{ L})(4 \text{ atm}) = (1.75 \text{ L}) P_2$$

$$P_2 = 0.56 \text{ atm}$$

# Charles And Boyles Law Gizmo Answer Key

**Michael Brown**

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## **Table of Contents Charles And Boyles Law Gizmo Answer Key**

1. Understanding the eBook Charles And Boyles Law Gizmo Answer Key
  - The Rise of Digital Reading Charles And Boyles Law Gizmo Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Charles And Boyles Law Gizmo Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Charles And Boyles Law Gizmo Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Charles And Boyles Law Gizmo Answer Key
  - Personalized Recommendations
  - Charles And Boyles Law Gizmo Answer Key User Reviews and Ratings
  - Charles And Boyles Law Gizmo Answer Key and Bestseller Lists
5. Accessing Charles And Boyles Law Gizmo Answer Key Free and Paid eBooks
  - Charles And Boyles Law Gizmo Answer Key Public Domain eBooks
  - Charles And Boyles Law Gizmo Answer Key eBook Subscription Services
  - Charles And Boyles Law Gizmo Answer Key Budget-Friendly Options

6. Navigating Charles And Boyles Law Gizmo Answer Key eBook Formats
  - ePUB, PDF, MOBI, and More
  - Charles And Boyles Law Gizmo Answer Key Compatibility with Devices
  - Charles And Boyles Law Gizmo Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Charles And Boyles Law Gizmo Answer Key
  - Highlighting and Note-Taking Charles And Boyles Law Gizmo Answer Key
  - Interactive Elements Charles And Boyles Law Gizmo Answer Key
8. Staying Engaged with Charles And Boyles Law Gizmo Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Charles And Boyles Law Gizmo Answer Key
9. Balancing eBooks and Physical Books Charles And Boyles Law Gizmo Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Charles And Boyles Law Gizmo Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Charles And Boyles Law Gizmo Answer Key
  - Setting Reading Goals Charles And Boyles Law Gizmo Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Charles And Boyles Law Gizmo Answer Key
  - Fact-Checking eBook Content of Charles And Boyles Law Gizmo Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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