

## Chapter 6 Worksheet

1. A sample of a pure element has a mass of 1.000 g and occupies a volume of 1.000 cm<sup>3</sup>. Calculate the density of the element.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$\text{Density} = \frac{1.000 \text{ g}}{1.000 \text{ cm}^3} = 1.000 \text{ g/cm}^3$$

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

Given: Mass = 10.0 g

Given: Volume = 2.000 cm<sup>3</sup>

Missing: Density  
 $\text{Ca}, \text{H}_2, \text{O}_2$

2. A sample of a pure element has a mass of 1.000 g and occupies a volume of 1.000 cm<sup>3</sup>. Calculate the density of the element.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$\text{Density} = \frac{1.000 \text{ g}}{1.000 \text{ cm}^3} = 1.000 \text{ g/cm}^3$$

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

Given: Mass = 10.0 g  
Volume = 2.000 cm<sup>3</sup>

Missing: Density  
 $\text{Ca}, \text{H}_2, \text{O}_2$   
1. The element is a metal.  
2. The element is a nonmetal.  
3. The element is a noble gas.

# Chem 107 C Laboratory Manual Answer Key

**M Mark**

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