

Complete the table with the help of the above-mentioned initial concentration calculation process.

	Initial moles of Vitamin C	Initial concentration of Vitamin C	Initial moles of Iodide (I <sup>-</sup> )	Initial concentration of Iodide (I <sup>-</sup> )	Initial moles of H <sub>2</sub> O <sub>2</sub>	Initial concentration of H <sub>2</sub> O <sub>2</sub>
Trial 1a	0.0028 mol	0.114 M			0.0088 mol	0.088 M
Trial 2	0.0028 mol	0.114 M			0.0088 mol	0.088 M
Trial 3	0.0028 mol	0.114 M			0.0088 mol	0.088 M
Trial 1b					0.0044 mol	0.044 M
Trial 1c						

Show your calculations here.

$$[\text{Vitamin C}] = .5g \times \frac{1 \text{ mol}}{176.1g} = 0.0028 \text{ mol} / 0.025L = 0.114 M$$

$$[I^-] = 0.02 \times 3 = 0.06$$

$$0.02 \times 2.5 = 0.05$$

$$[H_2O_2] = 0.03 \times 10 = 0.3g, \frac{1 \text{ mol}}{34.01g} = 0.0088 \text{ mol} / 0.100L = 0.088 M$$

$$0.03 \times 5 = 0.15g, \frac{1 \text{ mol}}{34.01g} = 0.0044 \text{ mol} / 0.100L = 0.044 M$$

# Chem 1212 Lab Manual

**Jack Duff, Alan Gabrielli**



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