

Complete the table with the help of the above-mentioned initial concentration calculation process.

	Initial moles of Vitamin C	Initial concentration of Vitamin C	Initial moles of Iodide (I ⁻)	Initial concentration of Iodide (I ⁻)	Initial moles of H ₂ O ₂	Initial concentration of H ₂ O ₂
Trial 1a	0.0028 mol	0.114 M			0.0060 mol	0.060 M
Trial 2	0.0028 mol	0.114 M			0.0060 mol	0.060 M
Trial 3	0.0028 mol	0.114 M			0.0014 mol	0.044 M
Trial 1b						
Trial 1c						

Show your calculations here.

$$[\text{Vitamin C}] = \frac{5\text{g} \times \frac{1\text{mol}}{170.1\text{g}}}{0.075\text{L}} = 0.0628\text{ mol} / 0.075\text{L} = 0.114\text{ M}$$

$$[\text{I}^-] = 0.02 \times 3 = 0.06$$

$$0.02 \times 2.5 = 0.05$$

$$[\text{H}_2\text{O}_2] = 0.03 \times 10 = 0.3\text{g} \times \frac{1\text{mol}}{34.01\text{g}} = 0.0088\text{ mol} / 0.100\text{L} = 0.088\text{ M}$$

$$0.03 \times 5 = 0.15\text{g} \times \frac{1\text{mol}}{34.01\text{g}} = 0.0044\text{ mol} / 0.100\text{L} = 0.044\text{ M}$$

Chem 1212 Lab Manual

Jack Duff, Alan Gabrielli

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