

Chem II Final Exam Study Guide

Lecture Notes

Chapter 14 - Solutions

Mixtures:

- Dissolve vs dissociate: dissolve means to break apart in a solution, dissociate means to break the molecule up into ions in a solution
- Entropy (S) = a drive to disorder
 - Mixtures are more disordered than pure substances

Intermolecular forces:

- Definition: interaction between individual molecules
- To break IMFs = release energy (endothermic)
- To build IMFs = to absorb energy (endothermic)
- Characteristics of strong IMFs:
 - Higher boiling points, lower melting points, higher viscosity, lower VP
- 4 types of IMFs:
 - Dispersion: in everything. Dispersion is **stronger in larger molecules**
 - Dipole-dipole: occurs in polar molecules, higher ΔEN = higher d-d forces
 - H bonds: hydrogen with N, O, F, occurs in polar molecules
- Polar dissolves polar, non polar dissolves non polar,
 - Less nonpolar shrubbiness = more polar
- Increasing / decreasing volume in a solution = looser, weaker bonds
- Examples of solutions:
 - Ideal solution: breaks the same type of bonds that were there originally $\Delta V = 0$, $\Delta T = 0$, $\Delta H = 0$
 - Miscible non-ideal: break weaker, form stronger, $\Delta V = (-)$ decrease, $\Delta T = (+)$ increase, $\Delta H = (-)$ exothermic
 - Tends to be immiscible, non ideal: ex: oil + water, break stronger IMFs, form weaker, $\Delta V = (+)$ increase $\Delta T = (-)$ decrease, $\Delta H = (+)$ endothermic

Concentration units:

- m = molality = mol solu / Kg solv
- M = molarity mol solu / L soln
- % m/v = g solu / mL soln
- ppm/ppb/ppt = 1 g solu / 10^6 g soln, 10^9 for ppb, 10^{12} for ppt
- Mole fraction: $X_a = \frac{n_a}{n_a + n_b}$

Henry's Law:

- Solubility of a gas when you're trying to dissolve it into a liquid
- $S = K_H \times P_{gas}$ K_H is given, pressure in ATM

Raoult's Law:

- **Use everything you see vapor pressure**
- Tendency for a mixture to evaporate

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