



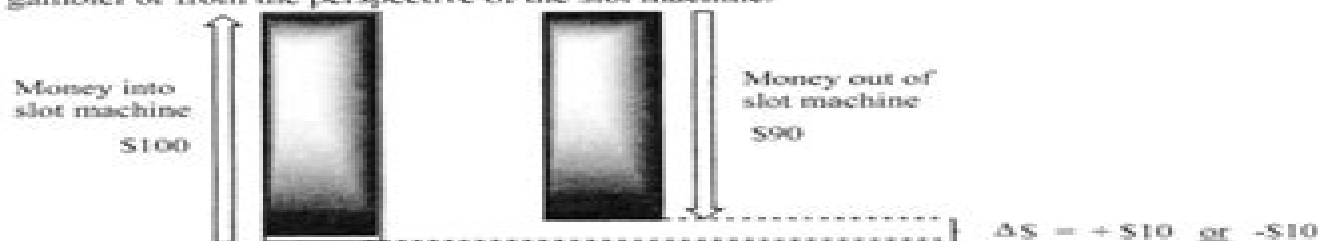
ChemActivity 6

Part A: Energy Diagrams

(What is the sign of $\Delta H_{\text{reaction}}$ for an exothermic reaction?)

Model 1A: The Slot Machine

Imagine a gambler sitting at a slot machine. Over the course of an hour the gambler puts \$100 into the machine and the machine pays out \$90. We know a net change (Δ) of ten dollars occurred, but what is the sign of this change in money (ΔS). To determine if $\Delta S = +10$ or -10 we need to decide if we calculate ΔS from the perspective of the gambler or from the perspective of the slot machine.



To translate this analogy into chemistry remember that:

- Money = energy.
- The slot machine = the reaction.
- The gambler = the experimenter (you).
- All changes (Δ) will be calculated from the perspective of the reaction (a.k.a. the slot machine).

Critical Thinking Questions

1. Using the convention listed above for calculating changes (Δ), what is the sign of ΔS in the example in Model 1A?
2. Based on this same convention, what sign must be attached to the \$100 that the gambler put into the slot machine in Model 1A? The sign of the \$90 that comes out is the opposite. **Add the appropriate sign to each value in Model 1A.**
3. On a different day the gambler puts in a total of \$20 and the slot machine pays out a total of \$100. What is the sign and value of ΔS for this day?



Chemactivity 6 Alien

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Table of Contents Chemactivity 6 Alien

1. Understanding the eBook Chemactivity 6 Alien
 - The Rise of Digital Reading Chemactivity 6 Alien
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemactivity 6 Alien
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemactivity 6 Alien
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemactivity 6 Alien
 - Personalized Recommendations
 - Chemactivity 6 Alien User Reviews and Ratings
 - Chemactivity 6 Alien and Bestseller Lists
5. Accessing Chemactivity 6 Alien Free and Paid eBooks
 - Chemactivity 6 Alien Public Domain eBooks
 - Chemactivity 6 Alien eBook Subscription Services
 - Chemactivity 6 Alien Budget-Friendly Options

6. Navigating Chemactivity 6 Alien eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemactivity 6 Alien Compatibility with Devices
 - Chemactivity 6 Alien Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemactivity 6 Alien
 - Highlighting and Note-Taking Chemactivity 6 Alien
 - Interactive Elements Chemactivity 6 Alien
8. Staying Engaged with Chemactivity 6 Alien
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemactivity 6 Alien
9. Balancing eBooks and Physical Books Chemactivity 6 Alien
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemactivity 6 Alien
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemactivity 6 Alien
 - Setting Reading Goals Chemactivity 6 Alien
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemactivity 6 Alien
 - Fact-Checking eBook Content of Chemactivity 6 Alien
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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