

## 16 • Chemical Equilibrium

1. Consider the reaction system.  
 $\text{CoO(s)} + \text{H}_2\text{g} \rightleftharpoons \text{Co(s)} + \text{H}_2\text{O(g)}$ .

The equilibrium constant expression is

a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$       d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$   
 b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$       c)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$   
 (d)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$   
 (d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$

2. Given the equilibrium,  
 $2\text{SO}_2\text{g} + \text{O}_2\text{g} \rightleftharpoons 2\text{SO}_3\text{g}$ , if this equilibrium is established by beginning with equal number of moles of  $\text{SO}_2$  and  $\text{O}_2$  in a 1.0 Liter bulb, then the following *must* be true at equilibrium:  
 a)  $[\text{SO}_2] = [\text{SO}_3]$       d)  $[\text{SO}_2] < [\text{O}_2]$   
 (d)  $2[\text{SO}_2] = 2[\text{SO}_3]$       e)  $[\text{SO}_2] > [\text{O}_2]$   
 c)  $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole  $\text{NO}$ , 0.200 mol  $\text{Cl}_2$  and 0.500 mol  $\text{ClNO}$  were placed in a 25.0 Liter container. The following equilibrium is established:  $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{g}$

3. At equilibrium, 0.600 mol of  $\text{ClNO}$  was present. The number of *moles* of  $\text{Cl}_2$  present at equilibrium is  
 (a) 0.050      (d) 0.200  
 b) 0.100      e) 0.250  
 c) 0.150

## PRACTICE TEST

4. The equilibrium constant,  $K_c$ , is:  
 a)  $4.45 \times 10^{-4}$       d) 0.167  
 b)  $6.67 \times 10^{-4}$       e) 1500  
 (c) 0.111

5. At  $985^\circ\text{C}$ , the equilibrium constant for the reaction,  
 $\text{H}_2\text{g} + \text{CO}_2\text{g} \rightleftharpoons \text{H}_2\text{O(g)} + \text{CO(g)}$

is 1.63. What is the equilibrium constant for the reverse reaction?  
 a) 1.63      d) 0.613  
 b) 0.815      (c) 1.00  
 c) 2.66

6. What is the relationship between  $K_p$  and  $K_c$  for the reaction,  $2\text{ICl(g)} \rightleftharpoons \text{I}_2\text{g} + \text{Cl}_2\text{g}$ ?  
 (a)  $K_p = K_c(RT)^{-1}$       d)  $K_p = K_c$   
 b)  $K_p = K_c(RT)$       e)  $K_p = K_c(2RT)$   
 c)  $K_p = K_c(RT)^2$

7. For the reaction  $2\text{NO}_2\text{g} \rightleftharpoons \text{N}_2\text{O}_4\text{g}$ ,  $K_p$  at  $25^\circ\text{C}$  is 7.3, when all partial pressures are expressed in atmospheres. What is  $K_c$  for this reaction? [ $R=0.0821 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ ]  
 (a) 4270      d) 179  
 b) 0.0119      e) 2.06  
 c) 0.291

# Chemical Equilibrium Test

**Crandall Zachariah Rosecrans, George  
Theodore Felbeck**

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Chemistry Joseph Albert Babor, Alexander Lehrman, 1929

*A Laboratory Manual in General Chemistry* William Foster, 1923

*Illustrated Guide to Home Chemistry Experiments* Robert Bruce Thompson, 2012-02-17

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