

LOHS Honors Chemistry

Balancing Chemical Equations

Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.

- $\text{BaCl}_2 + 2\text{KClO}_3 \rightarrow \text{Ba}(\text{ClO}_3)_2 + 2\text{KCl}$ DR
- $\text{Pb}(\text{OH})_2 \rightarrow \text{PbO} + \text{H}_2\text{O}$ D
- $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ SR
- $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2\text{KNO}_3$ DR
- $\text{Al}_2(\text{SO}_4)_3 + 3\text{Ca}(\text{OH})_2 \rightarrow 2\text{Al}(\text{OH})_3 + 3\text{CaSO}_4$ DR
- $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ SR
- $\text{Sn} + 2\text{KOH} \rightarrow \text{K}_2\text{SnO}_2 + \text{H}_2$ SR
- $2\text{PaI}_5 \rightarrow 2\text{Pa} + 5\text{I}_2$ D
- $3\text{SiF}_4 + 3\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SiF}_6 + \text{H}_2\text{SiO}_3$ SR
- $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_4$ S
- $16\text{Na} + \text{S}_8 \rightarrow 8\text{Na}_2\text{S}$ S
- $\text{N}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_2$ S

Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure include physical states where appropriate.

- Iron metal combines with oxygen gas to form iron (III) oxide solid



- Chromium metal and sulfuric acid (H_2SO_4) react to form chromium (III) sulfate solution and hydrogen gas.



- Barium chromate solution and aqueous hydrochloric acid (HCl) react to produce barium chloride solution and hydrogen chromate (solid).



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AN Whitehead



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