

Chemistry Benchmark II Study Guide

1. Write the formula for the following:
  - iron (II) oxide
  - calcium phosphate
  - ammonium phosphate
  - copper (III) bromide
  - nitrogen dioxide
  - titanium (IV) oxide
  - potassium bromide
  - magnesium hydroxide
  - xenon hexafluoride
2. What is a binary compound? What is a ternary compound?
3. Write formulas for:
  - hydroiodic acid
  - carbonic acid
  - hydrosulfuric acid
  - hydrochloric acid
  - hypochlorous acid
  - perchloric acid
  - phosphorous acid
  - chloric acid
  - sulfuric acid
  - chlorous acid
4. Name these:
  - NaCl
  - KI
  - CaBr
  - Rb<sub>2</sub>O
  - Sr<sub>2</sub>
  - CuCl
  - HgO
  - Fe<sub>2</sub>O<sub>3</sub>
  - MnO<sub>2</sub>
  - PbCl<sub>4</sub>
  - SnBr<sub>4</sub>
  - BF<sub>3</sub>
  - NO
  - N<sub>2</sub>O<sub>5</sub>
  - CCl<sub>4</sub>
  - P<sub>2</sub>O<sub>5</sub>
  - N<sub>2</sub>O<sub>3</sub>
5. Find the oxidation number assigned to:
  - S in H<sub>2</sub>SO<sub>4</sub>
  - P in PCl<sub>5</sub>
  - S in SO<sub>3</sub>
6. Name these:
  - NO<sub>2</sub>
  - N<sub>2</sub>O
  - SO<sub>3</sub>
7. How many valence electrons do elements have in the following groups?
  - alkali metals
  - alkaline earth metals
  - halogens
  - noble gases
8. Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
  - Sr
  - Br
  - Na
  - K
  - N
  - Cl
9. In each group, identify the atom with the largest atomic radius.
  - Sn, Xe, Rb, Sr
  - IRn, He, Xe, Kr
  - Pb, Ba, Cs, At
10. In each group, identify the atom with the highest ionization energy.
  - Cs, K, Li
  - Ba, Sr, Ca
  - I, Br, Cl
  - Mg, Si, S
11. Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond.
12. Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
13. Identify the type of bond in each:
  - K-N
  - Cs-O
  - C-N
  - O-F
  - Na-Cl
  - Cl-F
  - Cl-O
  - Ca-Cl
14. Who first arranged the periodic table in order of increasing atomic mass?
15. Draw Lewis structures for the following and predict the molecular shape.
  - H<sub>2</sub>S
  - PH<sub>3</sub>
  - O<sub>2</sub>
  - sulfate ion
  - SO<sub>3</sub>
  - BF<sub>3</sub>
16. Calculate the wavelength and energy of 96.6MHz radio waves.
17. What is the frequency and energy of electromagnetic radiation having a wavelength of  $3.33 \times 10^{-9}$ m?
18. Predict the ions that the following elements might form:
  - F
  - K
  - I
  - Ge
  - Fe (trick question)
  - P
  - Ag

# Chemistry Benchmark Study Guide Answer Key

**Betty A. Forbes, Daniel F. Sahm, Alice  
S. Weissfeld**

## Chemistry Benchmark Study Guide Answer Key:

*Study Guide for Bailey and Scott's Diagnostic Microbiology - E-Book* Betty A. Forbes, Daniel F. Sahm, Alice S. Weissfeld, 2016-06-30 Corresponding to chapters in Bailey Scott's Diagnostic Microbiology 12th Edition this new guide reviews important topics and helps students master key material. It includes chapter objectives a summary of key points review questions and case studies. Material is presented in an engaging format that challenges students to apply their knowledge to real life scenarios. Type Source Promotion Chapter Objectives open each chapter providing a measurable outcome to achieve by completing the material. A summary of Key Points from the main text helps students clearly identify key concepts covered in each chapter. Review Questions in each chapter test students on important knowledge in addition to key terms and abbreviations. Case studies in each chapter offer challenging questions for further analysis and challenge students to apply their knowledge to the real world.

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**The Publishers' Trade List Annual**, 1986 **Technical Book Review Index**, 1976 **Resources in Education**, 1997 **Bibliography of Agriculture**, 1992-11 **Subject Guide to Books in Print**, 1993 **Previews**, 1979 **Plating and Surface Finishing**, 1998 **El-Hi Textbooks and Serials in Print**, 1995 R R Bowker Publishing, 1995-02

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