



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride SF_6
- Lithium nitride Li_3N
- Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride SnCl_2
- Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- Mercury (I) chloride Hg_2Cl_2
- Potassium bromate KBrO_3
- Hydrosulfuric acid $\text{H}_2\text{S}(\text{aq})$
- Chloric acid HClO_3
- Sulfurous acid H_2SO_3

Name each of the following compounds.

- CuSO_4 Copper (II) sulfate
- AlF_3 aluminum fluoride
- HI hydrogen iodide
- NO nitrogen monoxide
- H_2Se hydroseonic acid
- HNO_3 nitric acid
- HNO_2 nitrous acid
- NaHSO_4 sodium bisulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives } 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

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Amin Elsersawi, Ph.D.

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Table of Contents Chemistry Answers

1. Understanding the eBook Chemistry Answers
 - The Rise of Digital Reading Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Answers
 - Personalized Recommendations
 - Chemistry Answers User Reviews and Ratings
 - Chemistry Answers and Bestseller Lists

5. Accessing Chemistry Answers Free and Paid eBooks
 - Chemistry Answers Public Domain eBooks
 - Chemistry Answers eBook Subscription Services
 - Chemistry Answers Budget-Friendly Options
6. Navigating Chemistry Answers eBook Formats
 - ePUB, PDF, MOBI, and More
 - Chemistry Answers Compatibility with Devices
 - Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Answers
 - Highlighting and Note-Taking Chemistry Answers
 - Interactive Elements Chemistry Answers
8. Staying Engaged with Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Answers
9. Balancing eBooks and Physical Books Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Answers
 - Setting Reading Goals Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Answers
 - Fact-Checking eBook Content of Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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