

## General Chemistry II Chapter 15 Study Guide

- $S_{\text{gas}} > S_{\text{liquid}} > S_{\text{solid}}$
- For similar molecules,  $S^\circ$  tends to increase with:
  - Increasing molar mass
  - Increased number of atoms in the formula

Standard entropy of a reaction ( $\Delta S_{\text{rxn}}^\circ$ ):

- $\Delta S_{\text{rxn}}^\circ$ : is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S_{\text{rxn}}^\circ = \sum m S_{\text{products}}^\circ - \sum n S_{\text{reactants}}^\circ$$

where  $m$  and  $n$  are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction:  $a \text{ A} + b \text{ B} \rightarrow c \text{ C} + d \text{ D}$

$$\sum m S_{\text{products}}^\circ = c \Delta S^\circ (\text{C}) + d \Delta S^\circ (\text{D})$$

$$\sum n S_{\text{reactants}}^\circ = a \Delta S^\circ (\text{A}) + b \Delta S^\circ (\text{B})$$

Remember to multiply  $\Delta H^\circ$  of each substance by its coefficient.

Predicting the sign of  $\Delta S_{\text{rxn}}^\circ$ :

- We can often predict the sign of  $\Delta S_{\text{rxn}}^\circ$  for processes that involve a change in the number of moles of gas
  - $\Delta S_{\text{rxn}}^\circ$  is **positive** if the amount of gas increases
  - $\Delta S_{\text{rxn}}^\circ$  is **negative** if the amount of gas decreases
  - $\Delta S_{\text{rxn}}^\circ$  is **likely to be positive** if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy ( $G$ ): combines the enthalpy and entropy of a system:
  - $G = H - TS$
- Gibbs free energy change ( $\Delta G$ ): is a measure of the spontaneity of a process and of the useful energy available from it:
  - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$  for a spontaneous reaction ( $-\Delta G$ )
- $\Delta G > 0$  for a nonspontaneous reaction ( $+\Delta G$ )
- $\Delta G = 0$  for a process at equilibrium ( $\Delta G = 0$ )

# Chemistry Chapter 15 Study Guide Answers

**Kenneth W. Raymond**



## **Chemistry Chapter 15 Study Guide Answers:**

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### **Table of Contents Chemistry Chapter 15 Study Guide Answers**

1. Understanding the eBook Chemistry Chapter 15 Study Guide Answers
  - The Rise of Digital Reading Chemistry Chapter 15 Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 15 Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 15 Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 15 Study Guide Answers
  - Personalized Recommendations
  - Chemistry Chapter 15 Study Guide Answers User Reviews and Ratings

- Chemistry Chapter 15 Study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 15 Study Guide Answers Free and Paid eBooks
  - Chemistry Chapter 15 Study Guide Answers Public Domain eBooks
  - Chemistry Chapter 15 Study Guide Answers eBook Subscription Services
  - Chemistry Chapter 15 Study Guide Answers Budget-Friendly Options
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  - Chemistry Chapter 15 Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 15 Study Guide Answers
  - Highlighting and Note-Taking Chemistry Chapter 15 Study Guide Answers
  - Interactive Elements Chemistry Chapter 15 Study Guide Answers
- 8. Staying Engaged with Chemistry Chapter 15 Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 15 Study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 15 Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 15 Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 15 Study Guide Answers
  - Setting Reading Goals Chemistry Chapter 15 Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 15 Study Guide Answers
  - Fact-Checking eBook Content of Chemistry Chapter 15 Study Guide Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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