

Ice Cream Lab

Pre-Lab Questions

1. What are the materials needed for this lab?
2. What are the reasons that ice cream freezes?
3. What 3 properties are involved in making ice cream? **WHY?**
4. How is ice cream made?
5. What is the best ice cream? Choose a flavor and sundae, toppings, etc. & **Why?**

DO try this at home!

The traditional method of ice cream making, using ice and rock salt to freeze a milk and sugar mixture, involves the application of several chemistry principles:

Colligative Properties: Physical properties determined by the concentration of dissolved particles in a mixture, and not affected by the type of dissolved particles. Osmotic pressure (tendency of water to flow from high water concentration to low water concentration as seen when applying salt to a slug), boiling point elevation and, most importantly in this activity, freezing point depression are examples of colligative properties. When 1 mole of NaCl dissolves in a liter of water it dissociates into 1 mole of sodium ions and 1 mole of chloride ions, or a total of 2 moles of particles per liter of mixture (a 2M concentration). When 1 mole of sugar dissolves in a liter of water there are just 1 mole of sugar molecules per liter (a 1M concentration). The greater the particle concentration, the stronger the colligative effect.

Ice cream, a mixture of milk, sugar and vanilla, is really an aqueous mixture with many particles in each liter. Because of this the freezing point of ice cream is lower than that of water, so to freeze ice cream, or to keep it frozen, you must keep its temperature significantly below 0 degrees Celsius. Adding rock salt to ice produces a melting ice and saltwater mixture with a depressed freezing point in which the ice cream can be frozen.

Heat of Fusion: Melting a solid to form a liquid is an endothermic process. The heat of fusion of a substance is the amount of energy needed to change 1 gram of solid to liquid. The heat of fusion of water is 334 joules/gram.

Since the room temperature and the milk mixture are both warmer than the freezing point of the ice mixture, energy will be transferred into the ice mixture on all sides. This energy will be used in melting the ice (334 joules per gram melted) and the temperature of the ice mixture will stay at its freezing point as long as there is still ice to melt. This process will effectively draw energy out of the milk mixture, lowering its temperature and, eventually, freezing it!

Hess' Law and Enthalpy Change: The energy absorbed by the ice mixture is transferred from the room and the milk mixture so $\Delta H_{\text{ice}} = \Delta H_{\text{milk mix}} + \Delta H_{\text{room}}$. This is an application of Hess' Law.

Chemistry Ice Cream Lab Questions

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Chemistry Ice Cream Lab Questions:

International Handbook of Research on STEAM Curriculum and Practice Stephen J. Farenga, Salvatore G. Garofalo, Daniel Ness, 2025-10-24 This comprehensive handbook delves into curriculum praxis human development and cognition within the contexts of the STEAM disciplines science technology engineering arts architecture and mathematics Cutting edge research will help educators identify best practice techniques for developing students knowledge in STEAM subjects as well as capture contemporary social and political issues within the STEAM context Drawing on the work of over 50 international contributors this volume covers both emergent and established areas of research giving voice to newcomers to the field as well as perspectives from established experts These areas are divided into five sections on foundations content teaching and learning throughout the lifespan equity and enrichment and settings Each topic is considered in both its historical and current context with a focus on the interconnections between theory and practice This book offers a first of its kind overview of STEAM curriculum development which will be especially useful to educational practitioners and researchers of STEAM subjects as well as teacher educators overseeing STEAM education This resource will also be useful for K 12 school and institutional libraries as reference material and for curriculum specialists and administrators seeking to identify methods of best educational practices within STEAM

The Use of Food in Chemistry Experiments to Engage and Enrich the Teaching in the Classroom Brian Michael Topping, 2010

Environmental Chemistry in the Lab Ruth Ann Murphy, 2022-08-31 Environmental Chemistry in the Lab presents a comprehensive approach to modern environmental chemistry laboratory instruction together with a complete experimental experience The laboratory experiments have an introduction for the students to read a pre lab for them to complete before coming to the lab a data sheet to complete during the lab and a post lab which would give them an opportunity to reinforce their understanding of the experiment completed Instructor resources include a list of all equipment and supplies needed for 24 students a lab preparation guide an answer key to all pre lab and post lab questions sample data for remote learners and a suggested rubric for grading the labs Additional features include Tested laboratory exercises with instructor resources for environmental science students Environmental calculations industrial regulation and environmental stewardship Classroom and remote exercises An excellent user friendly and thought provoking presentation which will appeal to students with little or no science background A qualitative approach to the chemistry behind many of our environmental issues today

Laboratory Assessment in the Chemistry Classroom Thomas Paul Hunt, 1998

The Complete Home Learning Sourcebook Rebecca Rupp, 1998 Lists all the resources needed to create a balanced curriculum for homeschooling from preschool to high school level

Applying Food and Other Household Materials to Beginning Chemistry Experiments Deborah Ann Pudell, 2000

Ice Cream Trade Journal , 1920 **The Ice Cream Trade Journal** , 1919 **Countertop Chemistry Experiment 15: Ice Cream** , The Science House at North Carolina State University in Raleigh presents Ice Cream as part of Countertop

Chemistry Countertop Chemistry is a collection of K 12 chemistry activities involving the use of chemicals found at home or in the hardware store For this experiment the students make ice cream in order to investigate how to change the freezing point of a solvent The Science House lists the materials needed highlights the procedures for the experiment provides notes for teachers and includes a list of related questions and answers The Science of Cooking Joseph J. Provost, Keri L. Colabroy, Brenda S. Kelly, Mark A. Wallert, 2016-05-02 The Science of Cooking The first textbook that teaches biology and chemistry through the enjoyable and rewarding means of cooking The Science of Cooking is a textbook designed for nonscience majors or liberal studies science courses that covers a range of scientific principles of food cooking and the science of taste and smell It is accompanied by a companion website for students and adopting faculty It details over 30 guided inquiry activities covering science basics and food focused topics and also includes a series of laboratory experiments that can be conducted in a traditional laboratory format experiments that can be conducted in a large class format and take home experiments that can be completed with minimal equipment at the student s home Examples of these engaging and applicable experiments include fermentation cheese and ice cream making baking the best cookies how to brown food faster and analyzing food components They are especially useful as a tool for teaching hypothesis design and the scientific process The early chapters of the text serve as an introduction to necessary biology and chemistry fundamentals such as molecular structure chemical bonding and cell theory while food based chapters cover Dairy products milk ice cream foams and cheeses Fruits and vegetables Meat and fish Bread Spices and herbs Beer and wine Chocolate and candies The Science of Cooking presents chemistry and biology concepts in an easy to understand way that demystifies many basic scientific principles For those interested in learning more science behind cooking this book delves into curious scientific applications and topics This unique approach offers an excellent way for chemistry biology or biochemistry departments to bring new students of all levels and majors into their classrooms General Catalog Iowa State University, 1911 *Gourmet Lab* Sarah Reeves Young, 2011 Hands on inquiry based and relevant to every student OCOs life Gourmet Lab serves up a full menu of activities for science teachers of grades 6 OCo 12 This collection of 15 hands on experiments OCO each of which includes a full set of both student and teacher pages OCO challenges students to take on the role of scientist and chef as they boil bake and toast their way to better understanding of science concepts from chemistry biology and physics By cooking edible items such as pancakes and butterscotch students have the opportunity to learn about physical changes in states of matter acids and bases biochemistry and molecular structure The Teacher pages include Standards addressed in each lab a vocabulary list safety protocols materials required procedures data analysis student questions answer key and conclusions and connections to spur wrap up class discussions Cross curricular notes are also included to highlight the lesson OCOs connection to subjects such as math and literacy Finally optional extensions for both middle school and high school levels detail how to explore each concept further What better topic than food to engage students to explore science in the natural world *Science Fair*

Project Index, 1985-1989 Cynthia Bishop, Katherine Ertle, Karen Zeleznik, 1992-06 Includes science projects and experiments found in 195 books published between 1985 and 1989 Almost all areas of science and many areas of technology are covered

The American Food Journal ,1921 Baker's Digest ,1947 *Industrial & Engineering Chemistry* ,1924

Chemical Abstracts ,1911 **Morphine, Ice Cream, Tears** Joseph Sacco, 1989 As a young idealistic resident doctor Joe Sacco witnessed helpless patients fall victims to the high tech medical stars Outraged he began to write about his experiences His story is one no reader will forget **Illinois Chemist** ,1921 **International Milk Dealer** ,1929

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