

Basics of Science (Ch 1)

- What are the steps in the scientific method?  
1-make observations      5-Experiment  
2-Ask a question/problem      6-Draw a conclusion  
3-Research      7-Theory  
4-Form a hypothesis
- Identify the following as qualitative or quantitative data  
Ql orange      QN 25°C  
Ql soft      Ql cold  
QN 55 mph      QN 15 mL
- Why must there always be a control group in an experiment?  
The control group is to compare the experimental group.
- What type of graph should be used in each of the following situations:
  - Data given as percentages      CIRCLE
  - a comparison is being made between different quantities      BAR
  - the relationship between two different variables is being illustrated.      LINE
- The variable that is changed by the experimenter is called the Manipulated or Independent variable.  
This variable is graphed on the X axis. (MIX)

### Bylaws amendments (Ch. 2)

3. Fill in the following table with the correct information.

Measurement	Equipment	Units	Definition
Volume	Graduated cylinder	mL, cm <sup>3</sup>	Space occupied
Mass	Balance	grams	Amount of matter
Time	Stopwatch	Seconds	Interval between two occurrences
Length	Ruler, meter stick	Meter	Distance between two points
Temperature	thermometer	Celsius, Kelvin	Avg. particle motion (kinetic energy)

2. Identify each of the following as either a derived or base unit.  
Mass b, volume d, length b, density d  
Temperature b, time b, speed d, area d

3. Define accuracy and precision. Draw a target representing accuracy and precision.  
Accuracy-a measurement that is close to the accepted value  
Precision- reproducing the same measurement (measurements close to each other)

4. The following measurements were taken during lab when measuring the distance a marble will roll: 12.43 mm; 12.40 mm; 12.39 mm. The calculated distance the marble should have rolled is 15.90 mm. Are these measurements accurate and/or precise? Explain your answer. PRECISE but NOT ACCURATE

a. A sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of the water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample.  
 $2.7 \text{ g/mL} = \text{M}/8 \text{ mL}$   $\text{M}=21.6 \rightarrow 22 \text{ g}$

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