

Chapter 19 – Nuclear Chemistry

This packet contains background and explanatory information on Nuclear Chemistry in addition to practice handouts. The information should be read and used to complete all of the practice included. I will be collecting this packet and giving points for the completion of the contents. You may want to consult your textbook (Chapter 19) or online resources for addition explanations and examples if needed.

Chapter Objectives: The student will (be able to)...

- Compare and contrast **alpha**, beta, and gamma radiation.
- Identify alpha decay, beta decay, fission, and fusion reactions.
- Balance nuclear equations.
- Solve half-life calculations.

Part I: Radioactivity - What is it? <http://www.darvill.clara.net/nucrad/index.htm>

All substance are made of **atoms**. These have **electrons** (e) around the outside, and a **nucleus** in the middle. The nucleus consists of **protons** (p) and **neutrons** (n), and is extremely small. (Atoms are almost entirely made of empty space!)

In some types of atom, the **nucleus is unstable**, and will **decay** into a more stable atom. This **radioactive decay** is completely **spontaneous**. The energy that is released from the nucleus of the atom is **radiation**.

When an unstable nucleus **decays**, there are **three** ways that it can do so. It may give out:-

- an **alpha particle** (we use the symbol α)
- a **beta particle** (symbol β)
- a **gamma ray** (symbol γ)

Many radioactive substances **emit** α particles and β particles as well as γ rays. In fact, you won't find a pure γ source: anything that gives off γ rays will also give off α and/or β too.

A. Alpha Particles – More Information

- Alpha particles are made of **2 protons and 2 neutrons**.
- This means that they have a **charge of +2**, and a **mass of 4** (the mass is measured in "atomic mass units", where each proton & neutron = 1)

We can write them as ${}^4_2\alpha$, or, because they're the same

as a helium nucleus, ${}^4_2\text{He}$.

- Alpha particles are relatively **slow and heavy**.
- They have a **low penetrating power** - you can stop them with just a **sheet of paper**.
- Because they have a large charge, alpha particles **ionize** (pull electrons from) other atoms strongly.



Alpha particle:
2 protons
& 2 neutrons

Chemistry Nuclear Packet Answers

Hussin A.Rothana



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Photochemistry Oliver Schalk, Enrico Tapavicza, 2021-03-30 Photochemistry is an important part of both chemistry and biology and is of great practical significance for the development of sustainable sources of energy The mechanisms of photochemistry are far from trivial and far from understood There are limits to how well theory can describe the processes and how well experiments can resolve them This book aims to provide an overview of state of the art methods for both theoretical development and experimental techniques with a focus on ultrafast molecular processes and electronic excitation of organic molecules These fields are active and progress is being made carried by the increasing speed of computation and the development of new light sources most notably X ray sources at large facilities Alongside these two layers of theoretical development and experimental techniques is a third layer model building In this layer model building tries to find similarities in seemingly unrelated experimental results and deepen our general knowledge of photoinduced processes Often progress is made not by cutting edge techniques but rather by using well established techniques with a great variety of molecules this approach promises less glory but is just as important as the first two layers Examples mentioned in the text are the Woodward Hoffman rules and the dynamophore concept All three layers are crucial to push our knowledge further and eventually to use it for developing new and more advanced optical devices

5 Practice Sets for NDA/ NA Exam Disha Experts, 2021-08-01 Chemistry Problems Joseph F. Castka, 1962 *Resources in Education*, 1994-04 *Physical Chemistry* John S. Winn, 1995 This text presents physical chemistry as a coherent whole rather than a set of disjointed topics and shows how the subject relates to the rest of chemistry and physics It emphasizes physical models as well as mathematical techniques along with both rigorous and approximate order of magnitude problem solving Designed to progress beyond a numerical answer problems expose the physical significance of the situation and teach students how to pose a problem in the first place In addition modern molecular concepts currently unanswered problems in research experimental techniques and new directions in the field are introduced wherever appropriate An orderly progression of thermodynamics carefully builds students knowledge without covering too much too early on Chemical reaction thermodynamics is covered in Chapter 7 after the culmination of thermodynamics with advanced material in Chapter 10

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