

CHEMISTRY CHEAT SHEET

INTERMOLECULAR FORCES

INTERMOLECULAR FORCES CHEAT SHEET

Intermolecular forces are the forces of attraction and repulsion between molecules. These forces are responsible for the physical properties of matter.

The strength of these forces is proportional to the size of the molecule. The stronger the interaction, the more energy is required to separate the molecules.

There are four types of intermolecular forces: London dispersion forces, dipole-dipole interactions, hydrogen bonding, and ion-dipole interactions.

London dispersion forces are the weakest type of intermolecular force. They are caused by the temporary dipoles in molecules.

Dipole-dipole interactions are stronger than London dispersion forces. They are caused by the permanent dipoles in molecules.

Hydrogen bonding is the strongest type of intermolecular force. It is caused by the strong permanent dipole in hydrogen atoms.

Ion-dipole interactions are caused by the attraction between a positive ion and a negative dipole.

The following table summarizes the properties of each type of intermolecular force.

Intermolecular Force	Description	Strength
London dispersion forces	Molecules with permanent dipoles attract each other.	Very weak
Dipole-dipole interactions	Molecules with permanent dipoles attract each other.	Stronger than London dispersion forces
Hydrogen bonding	Molecules with strong permanent dipoles attract each other.	Strongest of all intermolecular forces
Ion-dipole interactions	Positive ions attract negative dipoles.	Stronger than hydrogen bonding

Intermolecular forces are responsible for the physical properties of matter. They are responsible for the melting point, boiling point, and solubility of molecules.

The following table summarizes the properties of each type of intermolecular force.

Intermolecular Force	Description	Strength
London dispersion forces	Molecules with permanent dipoles attract each other.	Very weak
Dipole-dipole interactions	Molecules with permanent dipoles attract each other.	Stronger than London dispersion forces
Hydrogen bonding	Molecules with strong permanent dipoles attract each other.	Strongest of all intermolecular forces

Chemistry Physics Forces Study Guide

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