

# Chemistry 104 Lab 4

## 2024

### PRELAB 4

#### Thermochemistry: Heats of Reactions

You must show the entire calculation to receive full marks. It is not sufficient to record the answer only. Remember to use significant figures and units. Note that you may need to find your answers in the Online open-source textbooks and resources or Laboratory Manual.

1. Calculate the Heat Capacity of the Calorimeter using the provided information. 25.0g of hot water at 49.7°C was added to a Styrofoam cup calorimeter containing 25.0g of cold water at 27.2°C. After mixing, the temperature of the water in the calorimeter was determined to be 38.1°C. (See page 60 of Lab Manual) (4 marks)
2. Calculate the expected change in temperature given the following data:  
7.5 g of  $\text{NH}_4\text{NO}_3$  dissolved in 100.0 mL of  $\text{H}_2\text{O}$ , the molar heat of solution of  $\text{NH}_4\text{NO}_3$  is 25.7 kJ/mol. The density of  $\text{H}_2\text{O}$  is 1.00 g/mL. (5 marks)
3. Read, Lab 4 up to and including page 64 in the Chemistry 104 Laboratory Manual and initial here to indicate you have read the material. (1 mark)

# Chemistry 104 And 17 Practice Answers

**Darrell Ebbing**



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