

# CHEM 120 FINAL EXAM

6. (TCO 6) A gas at a temperature of 95 degrees C occupies a volume of 165 mL. Assuming constant pressure, determine the volume at 25 degrees C. Show your work. (Points : 5)

Using Charles' Law,  $(V1/T1) = (V2/T2)$   
First, convert temperature to KELVIN (T1 = 95 + 273)  
Thus,  $T1 = 95 + 273 = 368$

We have  $V1$  (165 mL) &  $T2 = (25 + 273) = 298$ ,  
 $V2 = (V1 * T2) / T1 = (165 \text{ mL} * 298) / 368 = 133.6 \text{ mL}$

7. (TCO 6) A sample of helium gas occupies 1021 mL at 719 mmHg. For a gas sample at constant temperature, determine the volume of helium at 745 mmHg. Show your work. (Points : 5)

$$1021 \text{ mL} \cdot 719 \text{ mmHg} / 745 \text{ mmHg} = 985.36 \text{ mL} = 985 \text{ mL}$$

Using Boyle's law,  $P1V1 = P2V2$ . We have  $V1$  (1021 mL),  $P1$  (719 mmHg) and  $P2$  (745 mmHg).

8. (TCO 12) If one strand of a DNA double helix has the sequence T T A G C G A C G C, what is the sequence of the other DNA strand? (Points : 10)

A A T C G C T G C C

1. (TCO 8) 35.0 mL of 0.25 M NaOH is neutralized by 23.6 mL of an HCl solution. The molarity of the HCl solution is (show your work): (Points : 5)

moles solute = moles solute / liters solution  
0.25 M = moles NaOH / 0.035 L  
moles NaOH = 0.00875 moles NaOH

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2. (TCO 1) How many mL are in 3.5 pints? Show your work. (Points : 5)

3.5 pints is equivalent to 1656.116  
1 pint = 473.176 mL  
3.5 pints \* 473.176 mL = 1656.116 mL

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