

Ch. 9 Reactions in Aqueous Solutions

Molarity and Solution Stoichiometry

- Molarity (M) is used to describe solution concentration because it is very easy to determine the number of moles of solute
 - Solute is what gets dissolved in the solvent
 - Solvent is water (typically)
 - Solution = solute + solvent
- Molarity = (#mol of solute(n)) / (Volume(V) of solution(L))
 - #mol solute = M x V
 - n = M x V
- Four things to know about molarity
 - M x V = # mol of solute
 - Molarity can be used as a conversion factor in stoichiometry problems.
 - You can always replace "M" with (mol/L)
 - An Important Molarity shortcut:

→ this equation is usually by derivation.

$$M = \frac{\text{# mol Solute}}{\text{# L Sol'n}} \cdot \frac{10^{-3}}{10^{-3}}$$

$$M = \frac{\text{# mol Solute}}{\text{# mL Sol'n}}$$

- Solution by dilution:
 - $M_1V_1 = M_2V_2$ or M(concentrated)V(concentrated)=M(dilute)V(dilute)

Problems:

- What is the concentration of a solution of HCl that was prepared by diluting 5.00 mL of 12.0 M HCl to 150.0 mL?

$$\frac{M_1V_1}{V_2} = \frac{M_2V_2}{V_2} \quad \frac{M_1V_1}{V_2} = M_2 \quad \frac{(12.0M)(5.00mL)}{(150.0mL)} = 0.400M$$

- How many L of 0.100 M HCl are required to completely react with 13.0 g of $MgCO_3$ in the following reaction?



$$\#L \text{ HCl} = 13.0g \text{ } MgCO_3 \times \frac{1 \text{ mol } MgCO_3}{84.32g \text{ } MgCO_3} \times \frac{2 \text{ mol HCl}}{1 \text{ mol } MgCO_3} \times \frac{L}{0.100 \text{ mol HCl}} = 3.08L \text{ HCl}$$

Electrolytes

- Be sure to review:
 - Names, formulas, and charges of all ions.
 - These terms: solution, solvent, solute, aqueous solutions
- Light bulb demonstration.

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