

25. The number of σ bonds in $\text{N}\equiv\text{N}$ is

- a. 1
- b. 2
- c. 3
- d. 4

26. The elements in an ionic compound are held together by

- a. electrostatic forces of attraction.
- b. van der Waals forces
- c. the spin of paired electrons.
- d. the formation of hybrid orbitals.
- e. an electron pair.

27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode

- a. at which oxidation occurs.
- b. which attracts cations.
- c. at which electrons are supplied to the solution.
- d. at which reduction occurs.

28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?

- a. 47
- b. 62
- c. 93
- d. 186

29. In a galvanic (voltaic) cell in which the reaction is $\text{Cd} + \text{Cu}^{2+} \rightarrow \text{Cu} + \text{Cd}^{2+}$ and the ions are at unit concentration (activity), the cell potential is.



30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?

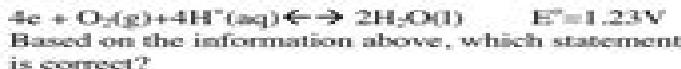
- a. $\text{CaCO}_3(s) \rightleftharpoons \text{CaO}(s) + \text{CO}_2(g)$
- b. $\text{H}_2(g) + \text{Cl}_2(g) \rightleftharpoons 2\text{HCl}(g)$
- c. $2\text{NO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$
- d. $\text{COCl}_2(g) \rightleftharpoons \text{CO}(g) + \text{Cl}_2(g)$

Standard Potentials	E°
$\text{Mg} \rightarrow \text{Mg}^{2+} + 2e^-$	2.37 V
$\text{Al} \rightarrow \text{Al}^{3+} + 3e^-$	1.66 V
$\text{Zn} \rightarrow \text{Zn}^{2+} + 2e^-$	0.76 V
$\text{Fe} \rightarrow \text{Fe}^{2+} + 2e^-$	0.44 V
$\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$	0.34 V
$\text{Ag} \rightarrow \text{Ag}^{+} + e^-$	-0.80 V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.

- a. Mg and Ag
- b. Mg and Fe
- c. Zn and Cu
- d. Al and Ag
- e. Mg and Al

32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
 $[\text{H}^+] = 1.0 \text{ M}$ initially, $P_{\text{O}_2} = 1.0 \text{ atm}$



- a. $n = 1$, since one mole of oxygen is being considered.
- b. Addition of base should result in an E value, which is less than 1.23 V.
- c. E is independent of the pH of the solution.
- d. $Q = \frac{[\text{H}_2\text{O}]}{[\text{O}_2][\text{H}^+]}$

33. The equilibrium constant for the gaseous reaction $\text{C} + \text{D} \rightleftharpoons \text{E} + \text{F}$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will

- a. proceed at equal rates in both directions.
- b. proceed more rapidly to form E and F.
- c. proceed more rapidly to form C and D.
- d. not occur in either direction.

Compound	$\Delta G^\circ_f, \text{kJ/mol}$
$\text{H}_2\text{O}(\text{l})$	-237
$\text{H}_2\text{O}(\text{g})$	-229

34. At 298 K the equilibrium constant for $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$

- a. is larger than the K_{eq} for $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g})$.
- b. will have a value of 1.0 at equilibrium.
- c. cannot be computed since data on O_2 and

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Ian Pickup

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