

Chemistry diagnostic test (Topic 12)

1. B – the p^2 config. of S is less stable, electron- electron repulsion.



2. A – B is lower than Be because two s electrons shield the 2p electrons
3. E – Au and Ag both have full d orbitals(10 electrons) (s^1)
4. E – electronegativity increases from lower left to upper right (F-most electronegative)
5. B – sulfur's oxidation numbers: 2-, 2+, 4+, 6+
6. A – atomic radii increase right to left and top to bottom
7. C – cations are smaller, the higher the charge the smaller the cation
8. A – Group 1 has low 1st ionization energy, but 2nd is pulling an inner core electron(big jump)
9. E – equal moles - 0.25
10. D – greatest molar mass
11. D – greatest mass/volume
12. E – all same temp
13. E – Phosphorus – high reactivity with oxygen not found free in nature
14. B – not all have a 2+ oxidation
15. A – all substances contain London dispersion forces
16. D – hydrogen must be bonded to nitrogen, oxygen or fluorine
17. A – propane (molar mass – 44) dispersion forces only – lowest temp
Ethanol (molar mass – 46) also has hydrogen bonding
Water (molar mass – 18) very strong hydrogen bonds
18. C
19. A – balanced equation is
 $Na_2CO_3 + 2HCl \rightarrow 2NaCl + CO_2 + H_2O$
Limiting is sodium carbonate
20. C – $(0.300 \text{ moles/L})(.300\text{L}) = 0.0900 \text{ mol Cl}^- \text{ ion}$ – requires 1/3 as much aluminum chloride or 0.0300 mol
 $0.0300 \text{ mol}/0.500\text{M} \text{ mol/mol/L} = 0.0600\text{L} = 60.0 \text{ ml}$
21. B - $0.500 \text{ M} \times 0.250\text{L} \times 200\text{g/mol} = 25.0 \text{ g}$
22. D- $2 C_3H_8O + 9O_2 \rightarrow 6CO_2 + 8H_2O$
23. C – 3.00 carbon = $\frac{1}{3}$ mole of C therefore $\frac{1}{3}$ mol of Z combines
 $0.5 \text{ mol Z} \times 36.0 \text{ g/mol}(3\text{timesC}) = 18.0 \text{ g Z}$
24. E – 4 moles of O
25. C – an alkali metal reacts with water to produce hydrogen gas and an aqueous alkali metal hydroxide
26. D – $H_2SO_4 + NaOH \rightarrow NaHSO_4 + H_2O$ (diprotonic acid and monobasic base)
27. E – X has a 3+ charge
28. E
29. B – 1 mole of magnesium requires 2 moles of HCl therefore HCl is limiting
30. D – a formation reaction produces 1 mole of a product

Chemistry Answer Key Diagnostic Test Topic 2

Mary Millhollon, Richard H. Langley

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