

**Chapter 17: Thermochemistry***Section 1: The Flow of Energy – Heat and Work*

1. What is thermochemistry? **The study of the energy changes that occur in changes of state or during chemical reactions.**
2. Heat always flows from a **warmer** object to a **cooler** object.
3. Given that  $1 \text{ cal} = 4.184 \text{ J}$ :
  - a. How many calories are in  $325.9 \text{ J}$ ? **77.89 cal**
  - b. How many J are in  $34.7 \text{ cal}$ ? **145.18 J**
4. What is the Law of Conservation of Energy (aka the First Law of Thermodynamics)? **States that in chemical and physical processes, energy is neither created nor destroyed – it is transferred or conserved.**
5. In an exothermic reaction:
  - a. Heat is **released** / absorbed
  - b.  $q$  is **positive** / **negative**
  - c.  $\Delta H$  is **positive** / **negative**
  - d. Heat is a **reactant** / **product**
6. In an endothermic reaction:
  - a. Heat is **released** / **absorbed**
  - b.  $q$  is **positive** / **negative**
  - c.  $\Delta H$  is **positive** / **negative**
  - d. Heat is a **reactant** / **product**
7. Diagram **A** represents an exothermic reaction, whereas diagram **B** represents an endothermic reaction:
8. What is the difference between heat capacity and specific heat capacity? **Specific heat is the amount of energy to increase the temperature of 1g of substance by  $1^\circ\text{C}$**
9. The temperature of a  $34.1 \text{ g}$  sample of a substance increases from  $29^\circ\text{C}$  to  $51^\circ\text{C}$  when the sample absorbs  $912 \text{ J}$  of heat. What is the specific heat of the substance?  **$1.126 \text{ J/g}^\circ\text{C}$**
10. When  $867 \text{ J}$  of heat are added to  $9.3 \text{ g}$  of a substance at STP, the temperature increases from  $0^\circ\text{C}$  to  $14^\circ\text{C}$ . What is the specific heat of the substance?  **$6.66 \text{ J/g}^\circ\text{C}$**

# Chemistry B Thermochemistry Packet Answers

**Lewis Morris**



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