

# Acids, Bases & pH Review - Honors Chemistry

Directions: For all multiple choice questions, highlight the correct answer(s). Answer short answer questions in the box provided.

## Part 1: Properties & Reactions of Acids & Bases

1. Complete the table below.

	ACIDS	BASES
Table		
pH range		
Conducting electricity		
Potassium		
LiOH		
Phenolphthalein		

2. An indicator:

- a. is the type of acid.
- b. is the type of base.
- c. has a neutral pH.
- d. changes color depending on the pH.

3. Litmus and phenolphthalein are indicators. What does that mean?

4. Acids and bases "dissociate" in water to form  $\text{H}^+$  (or  $\text{H}_3\text{O}^+$ ) and  $\text{OH}^-$  (or  $\text{H}_2\text{O}_2^-$ ). What does "dissociate" mean in this context?

5. Identify whether each substance is an acid or a base.

- a.  $\text{H}_2\text{O}$ :
- b.  $\text{HCl}$  (HCl):
- c.  $\text{Al(OH}_3\text{)}$ :
- d.  $\text{NH}_3 + \text{H}_2\text{O} \rightarrow \text{NH}_4^+ + \text{OH}^-$ ;  $\text{NH}_3$  is weak.
- e.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ ;  $\text{H}_2\text{SO}_4$  is strong.
- f.  $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$ ;  $\text{HNO}_3$  is strong.

# Chemistry Chapter 19 Acids Bases Study Guide Answers

**Susan McMurry, John McMurry**

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