



balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay

electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride SF_6

b. Lithium nitride Li_3N

c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$

d. Tin (II) chloride SnCl_2

e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)

f. Mercury (I) chloride Hg_2Cl_2

g. Potassium bromate KBrO_3

h. Hydrosulfuric acid $\text{H}_2\text{S(OH)}_2$

i. Chloric acid HClO_3

j. Sulfurous acid H_2SO_3

Name each of the following compounds.

a. CuSO_4 Copper (II) sulfate

b. AlF_3 aluminum fluoride

c. HI hydrogen iodide

d. NO nitrogen monoxide

e. H_2Se hydrogen selenide

f. HNO_3 nitric acid

g. HNO_2 nitrous acid

h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{r} 63 \times .692 = 43.596 \rightarrow 43 \\ 65 \times .308 = 20.02 \rightarrow 20 \\ \hline 63.62 \rightarrow 64 \end{array}$$

Chemistry Chapter 61 Assessment Answers

TD Snyder

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A Laboratory Manual of General Chemistry Leon Burr Richardson, Andrew Jackson Scarlett, 1929 [Pharmaceutical Chemistry E-Book](#) David G. Watson, 2011-02-09 This new book from the editor of the highly successful *Pharmaceutical Analysis* sets out to define the area of pharmaceutical chemistry as distinct from medicinal chemistry. It focuses less on prototypes of drugs that perhaps never came to market and more on the drugs currently in use. The emphasis in the book is on the physicochemical properties of drug molecules and in so far as they are known the way that these properties govern the interaction of the drug with its target. Important physicochemical properties include pKa and partition coefficient and the properties of the structural elements within the drug which provide interactions with the target via a range of intermolecular forces. The last fifteen years has seen a great advance in the knowledge of protein structures and a strong emphasis is given to the interaction of drugs with proteins which shape the majority of drug mechanisms. Features Focus on intramolecular actions. Mechanisms of action richly illustrated. Self assessment included. Comprehensive chapters on vitamins and biotechnological products.

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