

CHAPTER 15 **STUDY GUIDE****Energy and Chemical Change****Section 15.1 Energy**

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

1. Energy is the ability to do work or produce heat.
2. The law of conservation of energy states that energy *can be* created and destroyed.
3. Chemical potential energy is energy stored in a substance because of its *composition*.
4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
6. A *calorie* is the SI unit of heat and energy.
7. The *specific* heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.
8. *Kinetic* energy is energy of motion.
9. Chemicals participating in a chemical reaction contain only *potential energy*.
10. One nutritional Calorie is equal to *100 calories*.
11. One calorie equals *4.184 joules*.
12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Chemistry Energy Study Guide Key

Jane Aloi, Gina Erickson

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1. Understanding the eBook Chemistry Energy Study Guide Key
 - The Rise of Digital Reading Chemistry Energy Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Energy Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Energy Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Energy Study Guide Key
 - Personalized Recommendations
 - Chemistry Energy Study Guide Key User Reviews and Ratings
 - Chemistry Energy Study Guide Key and Bestseller Lists
5. Accessing Chemistry Energy Study Guide Key Free and Paid eBooks
 - Chemistry Energy Study Guide Key Public Domain eBooks
 - Chemistry Energy Study Guide Key eBook Subscription Services
 - Chemistry Energy Study Guide Key Budget-Friendly Options
6. Navigating Chemistry Energy Study Guide Key eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Energy Study Guide Key Compatibility with Devices
- Chemistry Energy Study Guide Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chemistry Energy Study Guide Key
- Highlighting and Note-Taking Chemistry Energy Study Guide Key
- Interactive Elements Chemistry Energy Study Guide Key

8. Staying Engaged with Chemistry Energy Study Guide Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemistry Energy Study Guide Key

9. Balancing eBooks and Physical Books Chemistry Energy Study Guide Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chemistry Energy Study Guide Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chemistry Energy Study Guide Key

- Setting Reading Goals Chemistry Energy Study Guide Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Energy Study Guide Key

- Fact-Checking eBook Content of Chemistry Energy Study Guide Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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