

CHAPTER 15 STUDY GUIDE

Energy and Chemical Change

Section 15.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- | | |
|-------|---|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat. |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed. |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> . |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object. |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius. |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy. |
| _____ | 7. The <i>specific heat</i> of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic energy</i> is energy of motion. |
| _____ | 9. Chemicals participating in a chemical reaction contain only <i>potential energy</i> . |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> . |
| _____ | 11. One calorie equals <i>4.184 joules</i> . |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat. |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule. |

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

Chemistry Energy Study Guide Key

Jane Aloï,Gina Erickson



Chemistry Energy Study Guide Key:

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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1. Understanding the eBook Chemistry Energy Study Guide Key
 - The Rise of Digital Reading Chemistry Energy Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Energy Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Energy Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Energy Study Guide Key
 - Personalized Recommendations
 - Chemistry Energy Study Guide Key User Reviews and Ratings
 - Chemistry Energy Study Guide Key and Bestseller Lists
5. Accessing Chemistry Energy Study Guide Key Free and Paid eBooks
 - Chemistry Energy Study Guide Key Public Domain eBooks
 - Chemistry Energy Study Guide Key eBook Subscription Services
 - Chemistry Energy Study Guide Key Budget-Friendly Options
6. Navigating Chemistry Energy Study Guide Key eBook Formats

- ePub, PDF, MOBI, and More
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- Chemistry Energy Study Guide Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Energy Study Guide Key
 - Highlighting and Note-Taking Chemistry Energy Study Guide Key
 - Interactive Elements Chemistry Energy Study Guide Key
- 8. Staying Engaged with Chemistry Energy Study Guide Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Energy Study Guide Key
- 9. Balancing eBooks and Physical Books Chemistry Energy Study Guide Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Energy Study Guide Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Energy Study Guide Key
 - Setting Reading Goals Chemistry Energy Study Guide Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Energy Study Guide Key
 - Fact-Checking eBook Content of Chemistry Energy Study Guide Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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