

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- $P_1V_1 = P_2V_2$

Charles' Theory and Boyle's Law

- Expansion of gases is caused by the activity of the gas molecules expanding the container.
- If the gas is compressible to small the volume it has large pressure exerts the pressure in the walls inward.
- The gas exerts pressure on walls of the container.
- Expansion of the gas will reduce.

Ex. 1. A balloon filled with hydrogen has a volume of 800 ml. at standard atmospheric pressure. After the balloon is released, it reaches an altitude of 8.0 km above the ground where the pressure is 0.080 atm. What is the volume of the balloon at this altitude?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.013 \times 10^5 \text{ Pa} \quad P_2 = 0.080 \times 1.013 \times 10^5 \text{ Pa} = 8.104 \times 10^3 \text{ Pa}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{800 \text{ ml} \times 1.013 \times 10^5 \text{ Pa}}{8.104 \times 10^3 \text{ Pa}} = 10000 \text{ ml}$$

Ex. 2. A gas at pressure of 1.013 x 10⁵ atm has a volume of 8.0 x 10⁻³ m³. If the pressure is increased until the volume is 0.001 m³, what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.013 \times 10^5 \text{ Pa} \quad P_2 = ? \quad V_1 = 8.0 \times 10^{-3} \text{ m}^3 \quad V_2 = 0.001 \text{ m}^3$$

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{1.013 \times 10^5 \text{ Pa} \times 8.0 \times 10^{-3} \text{ m}^3}{0.001 \text{ m}^3} = 8.104 \times 10^5 \text{ Pa}$$

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