

Solubility Rules:-NO₃⁻¹ all soluble-Cl⁻ all soluble except AgCl, Hg₂Cl₂ and PbCl₂-SO₄⁻² all soluble except BaSO₄, PbSO₄ and SrSO₄-OH⁻ all insoluble except Group 1, Ba(OH)₂, Sr(OH)₂, Ca(OH)₂-CO₃⁻² all insoluble except NH₄⁺ and Group 1-S⁻² all insoluble except Group 1 and Group 2 and NH₄⁺**Polyatomics:**-Acetate (C₂H₃O₂⁻²)-Chlorate (ClO₃⁻¹)-Chlorite (ClO₂⁻¹)-Cyanide (CN⁻¹)-Hydroxide (OH⁻¹)-Nitrate (NO₃⁻¹)-Nitrite (NO₂⁻¹)-Nitrite (NO₂⁻¹)-Carbonate (CO₃⁻²)-Chromate (CrO₄⁻²)-Dichromate (Cr₂O₇⁻²)-Oxalate (C₂O₄⁻²)-Sulfate (SO₄⁻²)-Sulfite (SO₃⁻²)-Arsenate (AsO₄⁻³)-Arsenite (AsO₃⁻³)-Phosphate (PO₄⁻³)-Phosphite (PO₃⁻³)-Ammonium (NH₄⁺¹)-Bicarbonate (HCO₃⁻¹)**Acids:**

-HCl = hydrochloric

-H₂SO₄ = sulfuric-HNO₃ = nitric-CH₃COOH = acetic-H₃PO₄ = phosphoric-H₂CO₃ = carbonicMatter is the physical material of the universe and exists in 3 states: solid, liquid, gasSubstance is matter that has a uniform and definite compositionMixture is a physical blend of substancesHomogenous Mixture has uniform composition throughout (solution) Heterogeneous Mixture does notPhysical Properties: observed or measuredChemical Properties: ability to undergo chemical changeLaw of Conservation of Mass: matter/mass can be neither created nor destroyedExtensive Properties: mass and volumeIntensive Properties: compositionSolid: definite shape, volume, not easily compressed Liquid: indefinite shape, definite volume, not easily compressedGas: indefinite shape, volume, easily compressed *Particles are always moving in S, L, GIndications of Chemical Reaction: dissolving, releasing/absorption of heat, production of gas, color change, precipitateQualitative give descriptions, no numbersQuantitative number-basedAccuracy is how close to a single measurementPrecision is how close several measurements are together

Error = |expt - accepted| % Error = |expt-accepted|/accepted * 100 Percent Yield = expt/theoretical * 100

1 kilo(unit) = 1000 (unit)

1 (unit) = 1000 milli

1 (unit) = 100 centi

micro=10⁶nano=10⁹

K = C + 273

C = K - 273

H₂O FP=273K

BP=373K Density = mass/volume

Density of H₂O = 1 g/mLSig Figs: For addition & subtraction use least # of decimal places

/ For Multiplication & Division use least # of sig figs

Dalton's Atomic Theory: 1) all elements are composed of atoms

2) atoms of same element are identical

3) atoms can physically mix/chemically combine in whole #d ratios

4) chemical rxns happen; atoms don't change

Proton = pos., change = radioactivityNeutron = neutral, change = isotopeElectron = neg., change = ion

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