

## Intro to Chemistry Final Exam Study Guide

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. The Fahrenheit temperature scale is defined by which of the following temperatures?
  - a. Water boils at 100 degrees and freezes at 32 degrees.
  - b. Water boils at 212 degrees and freezes at 32 degrees.
  - c. Water evaporates at 212 degrees and freezes at 0 degrees.
  - d. Liquid water turns to a gas at 100 degrees and to a solid at 0 degrees.
2. A cold-blooded reptile basks on a warm rock in the sun. Its body is warmed by
  - a. radiation.
  - b. conduction.
  - c. convection.
  - d. Both (a) and (b)
3. Temperature is a measure of the average \_\_\_\_\_ energy of an object's particles.
  - a. mechanical
  - b. kinetic
  - c. potential
  - d. light
4. The temperature at which all molecular motion stops is
  - a. 0°C.
  - b. 0°F.
  - c. 0 K.
  - d. All of the above
5. Convert 500°F to degrees Celsius.
  - a. 260°C
  - b. 290°C
  - c. 84.2°C
  - d. 95.8°C
6. Convert 300 K to the Celsius scale.
  - a. 0°C
  - b. 27°C
  - c. 73°C
  - d. 573°C
7. Convert 458 K to the equivalent Fahrenheit temperature.
  - a. 185°F
  - b. 237°F
  - c. 365°F
  - d. 731°F
8. Which of the following temperatures is impossible to measure?
  - a. -85°F
  - b. -50°C
  - c. -20 K
  - d. 545°F
9. What is -175°C on the Kelvin scale?
  - a. 76 K
  - b. 89 K
  - c. 98 K
  - d. 448 K
10. The principles of atomic theory recognized today were conceived by
  - a. Avogadro.
  - b. Bohr.
  - c. Dalton.
  - d. Rutherford.
11. The total number of protons and neutrons in the nucleus of an atom is its
  - a. atomic number.
  - b. Avogadro number.
  - c. mass number.
  - d. average atomic mass.
12. All atoms of the same element have the same
  - a. atomic mass.
  - b. number of neutrons.
  - c. mass number.
  - d. atomic number.
13. Atoms of the same element can differ in
  - a. chemical properties.
  - b. mass number.
  - c. atomic number.
  - d. number of protons and electrons.
14. Chlorine has atomic number 17 and mass number 35. It has
  - a. 17 protons, 17 electrons, and 18 neutrons.
  - b. 35 protons, 35 electrons, and 17 neutrons.
  - c. 17 protons, 17 electrons, and 52 neutrons.
  - d. 18 protons, 18 electrons, and 17 neutrons.
15. Phosphorus-33 (atomic number 15) contains
  - a. 33 protons.
  - b. 18 neutrons.
  - c. 33 neutrons.
  - d. 18 protons.
16. Neon-22 contains 12 neutrons. It also contains
  - a. 12 protons.
  - b. 22 protons.
  - c. 22 electrons.
  - d. 10 protons.
17. Valence electrons determine an atom's

# Chemistry Final Exam Study Guide Multiple Choice

**Chirag Amin**

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*Resources in Education*, 1983-04

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