

# Chemistry Finals Study Guide

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## Chapter 1 & 2 - Review of Basic Concepts

- **Describe the 3 states of matter.** The three states of matter are solid, liquid, and gas.
  - **Solid:** When atoms in a material are moving slow that they stick together and barely move, creating a hard object classified as a solid
  - **Liquid:** When atoms are moving slower than a gas state but faster than the solid state so they flow and are able to take the shape of their container. They have more energy than solid state atoms.
  - **Gas:** When atoms are moving faster than the other two states of matter (solid & liquid) that they are no longer pulled towards each other and roam freely without boundaries if not kept in a container.
- **Describe the difference between elements and compounds, mixtures.**
  - **Elements:** materials that are only made out of the same type of atoms and cannot be changed into simpler substances by a chemical change (ex: O<sub>2</sub>)
  - **Compounds:** materials that are a mix of two or more element in a fixed proportion
  - **Mixtures:** a mixture of 2 or more pure substances.
    - **Heterogeneous Mixture:** can often be separated by filtration.
    - **Homogeneous Mixture:** can be separated by distillation, crystallization, and chromatography
- **Convert between Celsius and Kelvin.**
  - **Celsius to Kelvin:**  $K = ({}^{\circ}C + 273.15)$
  - **Kelvin to Celsius:**  ${}^{\circ}C = (K - 273.15)$
- **Round measured and calculated values to the correct number of significant digits.**
  - Significant Digit Rules:  [CHEM] Final 1 - Steps & Processes
- **Use dimensional analysis to solve multi step equations, showing how units cancel**
  -  [CHEM] Final 1 - Steps & Processes
- **What is the difference between precision and accuracy?**
  - **Precision:** Acquiring the same result again and again
  - **Accuracy:** A result that is close to the accepted value/actual answer

# Chemistry Final Exam Study Guide Part

**Princeton Review (Firm)**

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