

Gases | Study Guide

Test Format

Multiple Choice

Gas Law Problems

Gas Stoich Problems

Short Answer

Equations & Conversions

$$P = \frac{F}{A}$$

$$1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{\text{Rate of Effusion of L}}{\text{Rate of Effusion of H}} = \frac{\sqrt{\text{Molar Mass of H}}}{\sqrt{\text{Molar Mass of L}}}$$

$$PV = nRT$$

$$K = {}^{\circ}\text{C} + 273 \quad 0.001 \text{ kL} = 1 \text{ L} = 1,000 \text{ mL}$$

$$R = 0.0821 \text{ L} \cdot \text{atm}/\text{mol} \cdot \text{K}$$

$$\text{At STP, 1 mole} = 22.4 \text{ L} \quad \text{KHDSdcm}$$

1. A box weighs 180 lbs and has a base that is 6 inches by 9 inches. What is the pressure this box exerts on the surface it is sitting on?

$$P = \frac{F}{A} \quad \frac{180 \text{ lbs}}{54 \text{ in}^2} = 3.3 \text{ psi}$$

$54 \text{ in}^2 = \text{area}$

2. What is atmospheric pressure?

The pressure exerted by the collisions of gases in the air surrounding Earth (N₂, O₂, etc.).

3. Convert 1.088 atm to:

a. mm Hg $\frac{1.088 \text{ atm}}{1} \cdot \frac{760 \text{ mmHg}}{1 \text{ atm}} = 826.9 \text{ mmHg}$

b. kPa $\frac{1.088 \text{ atm}}{1} \cdot \frac{101.3 \text{ kPa}}{1 \text{ atm}} = 110.2 \text{ kPa}$

4. What is a barometer and who invented it?

A tool that measures room/air pressure. Invented by Torricelli.



5. How are temperature and kinetic energy related?

Direct: As temp \uparrow , KE \uparrow , molecules move faster.
As temp \downarrow , KE \downarrow , molecules move slower.

6. Briefly describe the five assumptions of the kinetic-molecular theory.

Gases have a lot of tiny particles that are always in motion that is random, fast, and continuous. Particles are far apart without any forces of attraction, so their collisions are elastic \rightarrow bouncing off of each other/walls. The speed of the molecules change with temperature.

Chemistry Gases Study Guide Answers Teacher Guide

J Spring

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Table of Contents Chemistry Gases Study Guide Answers Teacher Guide

1. Understanding the eBook Chemistry Gases Study Guide Answers Teacher Guide
 - The Rise of Digital Reading Chemistry Gases Study Guide Answers Teacher Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Gases Study Guide Answers Teacher Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Gases Study Guide Answers Teacher Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Gases Study Guide Answers Teacher Guide
 - Personalized Recommendations
 - Chemistry Gases Study Guide Answers Teacher Guide User Reviews and Ratings
 - Chemistry Gases Study Guide Answers Teacher Guide and Bestseller Lists
5. Accessing Chemistry Gases Study Guide Answers Teacher Guide Free and Paid eBooks
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 - Chemistry Gases Study Guide Answers Teacher Guide eBook Subscription Services
 - Chemistry Gases Study Guide Answers Teacher Guide Budget-Friendly Options

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 - ePUB, PDF, MOBI, and More
 - Chemistry Gases Study Guide Answers Teacher Guide Compatibility with Devices
 - Chemistry Gases Study Guide Answers Teacher Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Gases Study Guide Answers Teacher Guide
 - Highlighting and Note-Taking Chemistry Gases Study Guide Answers Teacher Guide
 - Interactive Elements Chemistry Gases Study Guide Answers Teacher Guide
8. Staying Engaged with Chemistry Gases Study Guide Answers Teacher Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Gases Study Guide Answers Teacher Guide
9. Balancing eBooks and Physical Books Chemistry Gases Study Guide Answers Teacher Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Gases Study Guide Answers Teacher Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Gases Study Guide Answers Teacher Guide
 - Setting Reading Goals Chemistry Gases Study Guide Answers Teacher Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Gases Study Guide Answers Teacher Guide
 - Fact-Checking eBook Content of Chemistry Gases Study Guide Answers Teacher Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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