

**10****CHEMICAL QUANTITIES****Practice Problems***On a piece of paper, solve the following problems.****YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!*****SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

- What is the molar mass of sucrose ( $C_{12}H_{22}O_{11}$ )?
- What is the molar mass of each of the following compounds?
  - phosphorus pentachloride ( $PCl_5$ )
  - uranium hexafluoride ( $UF_6$ )
- Calculate the molar mass of each of the following ionic compounds:
  - $KMnO_4$
  - $Ca_3(PO_4)_2$
- How many moles is  $3.52 \times 10^{24}$  molecules of water?
- How many atoms of zinc are in 0.60 mol of zinc?
- What is the mass of 1.00 mol of oxygen ( $O_2$ )?

**SECTION 10.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS**

- What is the molar mass of each of the following compounds?
  - $C_6H_{12}O_6$
  - $NaHCO_3$
  - $C_2H_6$
  - $KNH_4SO_4$
- Calculate the mass in grams of each of the following:
  - 8.0 mol lead oxide ( $Pb_3O_4$ )
  - 1.50  $\times 10^{-2}$  mol molecular oxygen ( $O_2$ )
  - 0.75 mol hydrogen sulfide ( $H_2S$ )
  - 2.30 mol ethylene glycol ( $C_2H_6O_2$ )
  - 0.00100 mol silicon tetrahydride ( $SiH_4$ )
- How many grams are in 1.73 mol of dinitrogen pentoxide ( $N_2O_5$ )?
- How many grams are in 0.658 mol of calcium phosphate [ $Ca_3(PO_4)_2$ ]?
- Calculate the number of moles in each of the following:
  - 0.50 g sodium bromide ( $NaBr$ )
  - 0.00100 g monochloromethane ( $CH_3Cl$ )
  - 13.5 g magnesium nitrate [ $Mg_2(NO_3)_3$ ]
  - 1.50  $\times 10^{-3}$  g propylene glycol [ $C_3H_6(OH)_2$ ]
  - 1.02 g magnesium chloride ( $MgCl_2$ )
- A chemist plans to use 435.0 grams of ammonium nitrate ( $NH_4NO_3$ ) in a reaction. How many moles of the compound is this?
- A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $C_{20}H_{24}N_4O_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
- What is the volume at STP of 2.66 mol of methane ( $CH_4$ ) gas?
- How many moles is 13.5 L of ammonia ( $NH_3$ ) gas at STP?

# Chemistry Ionic Study Guide Answer Key

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