



## Reaction Rates and Equilibrium Study Guide

### ANSWER KEY

- Describe how each of the following factors would affect the rate of a chemical reaction in terms of "particle collisions" in your response.
  - Temperature decreases** *If temperature decreases, the particles move slower and there will be less reactant particle collisions AND a smaller percentage of particles have enough energy to produce a reaction. The reaction rate will therefore decrease.*
  - Addition of iodide ions,  $I^-$ , as a catalyst** *The presence of the catalyst speeds up the reaction taking place. This means that a larger percentage of particles have enough energy to produce a reaction. The reaction rate will therefore increase.*
  - Increased reactant concentration** *With more reactant particles present, there are more collisions between those particles. This means the reaction rate will increase.*
  - Crushing a solid into a powder (with larger surface area)** *As a powder, there is more surface area for collisions to take place. This means there will be more collisions and a faster reaction rate.*
- The reaction progress diagram shows the energy changes that take place when nitrogen dioxide gas reacts with carbon monoxide gas.

# Chemistry Reaction Rates And Equilibrium Study Guide

**Seager/Slabaugh**



## **Chemistry Reaction Rates And Equilibrium Study Guide:**

**Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice** Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 Barron's Science 360 Chemistry is your complete go to guide for everything chemistry This comprehensive guide is an essential resource for High school and college courses Homeschooling Virtual Learning Learning pods Inside you'll find Comprehensive Content Review Begin your study with the basic building block of chemistry and build as you go Topics include atomic structure chemical formulas electrochemistry the basics of organic chemistry and much more Effective Organization Topic organization and simple lesson formats break down the subject matter into manageable learning modules that help guide a successful study plan customized to your needs Clear Examples and Illustrations Easy to follow explanations hundreds of helpful illustrations and numerous step by step examples make this book ideal for self study and rapid learning Practice Exercises Each chapter ends with practice exercises designed to reinforce and extend key skills and concepts These checkup exercises along with the answers and solutions will help you assess your understanding and monitor your progress Access to Online Practice Take your learning online for 50 practice questions designed to test your knowledge with automated scoring to show you how far you have come **The Complete Chemistry Study Guide and Note Cards and MCAT** Konstantinos Papadopoulos, 2012-07-06 *Study Guide for E. Russell Hardwick's Introduction to Chemistry* Joan Bouillon, 1984 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Chemical Principles Study Guide/Solutions Manual* John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications **Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems **Student Study Guide Fundamentals of Chemistry by David E. Goldberg** Margaret G. Kimble, 1994 *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote

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