



# CHEMISTRY

#### PERIODIC TABLE OF THE ELEMENTS

PERIODIC TABLE OF THE ELEMENTS																	
1	H	2	He	3	Li	4	Be	5	B	6	C	7	N	8	O	9	F
11	Na	12	Mg	13	Al	14	Si	15	P	16	S	17	Cl	18	Ar	19	K
21	K	22	Ca	23	Sc	24	Ti	25	V	26	Cr	27	Mn	28	Fe	29	Co
31	Rb	32	Sr	33	Ta	34	W	35	Ru	36	Rh	37	Pt	38	Ir	39	Pd
41	Cs	42	Ba	43	Hf	44	Ta	45	Os	46	Os	47	Os	48	Os	49	Os
51	Fr	52	Ra	53	Hg	54	Tl	55	Pb	56	Pb	57	Pb	58	Pb	59	Pb
1	H	2	He	3	Li	4	Be	5	B	6	C	7	N	8	O	9	F
11	Na	12	Mg	13	Al	14	Si	15	P	16	S	17	Cl	18	Ar	19	K
21	K	22	Ca	23	Sc	24	Ti	25	V	26	Cr	27	Mn	28	Fe	29	Co
31	Rb	32	Sr	33	Ta	34	W	35	Ru	36	Rh	37	Pt	38	Ir	39	Pd
41	Cs	42	Ba	43	Hf	44	Ta	45	Os	46	Os	47	Os	48	Os	49	Os
51	Fr	52	Ra	53	Hg	54	Tl	55	Pb	56	Pb	57	Pb	58	Pb	59	Pb

#### GENERAL INFORMATION

• MATTER

1. Chemical changes involve the breaking and/or formation of chemical bonds. Physical changes of substances do not result in the formation of new substances.
2. In a reaction, matter is neither created nor destroyed, but only changed from one form to another.
3. Mass is a measure of the amount of matter in an object.

THE PRACTICE

1. Energy is the ability to do work or transfer heat.
2. Energy appears in such forms as light, sound, heat, or mechanical, electrical, or chemical energy.
3. Energy can be converted from one form to another, but can be neither created nor destroyed.
4. Two general classifications of energy are

**C. ONE MOLE OF ANY SUBSTANCE** – is that which contains Avogadro's number ( $6.02 \times 10^{23}$ ) of particles. Number of moles = mass/molecular mass.

## ATOMIC STRUCTURE AND THE PERIODIC TABLE

## **D. THE BOHR THEORY OF THE HYDROGEN ATOM**

1. Electrons can exist in only certain stable energy levels.
2. When the electronic state of an atom changes, it must absorb or emit exactly that amount of energy equal to the difference between the final and initial states.
3. The wavelengths of the lines in the spectrum of an atom are characteristic of the atom.

the percentage of hydrogen atom:

where  $\beta = 1.039,678 \text{ cm}^{-1}$ ,  $m_1$  and  $m_2$  are quantum numbers for states  $a$  and  $b$ , and  $\delta$  is the separation of field-modified or uncorrected

**THE WAVE MECHANICAL MODEL**  
1. A wave function corresponds to a certain electronic energy and describes a region about the nucleus (orbital) where an electron having that energy may be found.



BRUNNEN

— 100 —

1. An ionic bond occurs when one or more electrons are transferred from the valence shell of one atom to another and there is a coulombic attraction between oppositely charged ions. The atom that loses electrons becomes a positive ion (cation), and the atom that acquires electrons becomes a negatively charged ion (anion).
2. The octet rule states that atoms tend to gain or lose electrons until their valence shells are filled.
3. In a nonpolar covalent bond, the electrons are shared equally between atoms. In a polar covalent bond, there is unequal sharing.

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**Ray Lester**

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