

# Chemistry Semester 2 Final Exam Study Guide

Name \_\_\_\_\_

- ✓ The final exam will be all multiple-choice. Come prepared with a pencil.
- ✓ Do not bring electronic devices to the exam (other than a calculator). This includes cell phones and mp3 players. If I see them I will take them.
- ✓ There will be no rest room passes until you are done with the test.
- ✓ I will provide a formula sheet (identical to one attached to this study guide) and a laminated periodic table.
- ✓ This study guide will be extra credit toward your final exam grade.
- ✓ **BRING A PENCIL AND A CALCULATOR!!!** Don't come to my final without these.
- ✓ The Chemistry Final Exam will count for 15% of your final semester grade. The exam will cover material from second semester. Some questions will require you to know material from first semester. Complete this review sheet on a separate sheet of paper. This review sheet covers all of the questions you will answer on the final exam, but be prepared to apply the knowledge on this study guide to other questions.

"Success is a state of mind. If you want success, start thinking of yourself as a success."

- ✓ Keys to success:
  - Complete this study guide.
  - Don't wait until Monday or Tuesday before finals to get help.
  - Start studying early. Remember, you have all summer to relax.
  - Be prepared.

## Final exam topics:

1. Chemical Reactions + Balancing
2. The Mole + Stoichiometry
3. Aqueous solutions and Solubility
4. The Behavior of Gases
5. Acids and Bases

## Final Exam Schedule

Find your class period and time. DO NOT be late. You will not be compensated for time you are late and your name will be turned into the office. There is NO excuse to miss the final exam.

### Thursday, May 27

8:10-9:40	Period 3
9:50-11:20	Period 4
11:20-12:15	Lunch
12:20-1:50	Period 5
2:00-3:30	Period 7

### Friday, May 28

8:10-9:40	Period 1
9:50-11:20	Period 2
11:20-12:15	Lunch
12:20-1:50	Period 6
2:00-3:30	Period 8 Make ups

# Chemistry Semester 2 Final Study Guide Answers

**Oswaal Editorial Board**



## **Chemistry Semester 2 Final Study Guide Answers:**

Directory of Distance Learning Opportunities Modoc Press, Inc., 2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation

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Papers for total 15 Subjects including English I English II Physics Chemistry Biology History Civics Geography Mathematics Hindi Computer Application Economics Economic Applications Commercial Studies Commercial Applications and Physical Education

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**5 Steps to a 5 AP Chemistry, 2014-2015 Edition** Richard H. Langley, John Moore, 2013-08-02 A

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**Catalog of Carleton College for the Academic Year ...** Carleton College (Northfield, Minn.), 1912

**Proceedings** Bassam Z. Shakhashiri, 1973

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**Register of Vanderbilt University ... Announcement ...** Vanderbilt University, 1928

The Software Encyclopedia 2001, 2001

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web answer the following questions based on the solubility curve below which salt is least soluble in water at 20 c 2 how many grams of potassium chloride can be dissolved in 200 g of water at 80 c io 3 at 40 c how much potassium nitrate can be dissolved in 300 g of water w 4 which salt shows the least change

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web solvents the larger volume of the mixture and solutes the smaller volume of the mixture for example a hot chocolate is a solution in which the solute the chocolate powder is dissolved in the solvent the milk or water the solute and solvent can be

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web they should be able to come up with the following answers stir the water increase the temperature and increase the amount of water alternately you may challenge the students to a contest

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web use a solubility curve to determine whether a solution is saturated or unsaturated and to estimate the amount of excess solute that remains undissolved in the solution

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web give results of all calculations in blue cells you must put your answers in the indicated cells as the grade book pulls your data from those cells do not put any units in the number cells use the unit boxes given graph tab transfer from data page the temp and concentration of salt in units of g salt 100 g water

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web 30 09 99 solubility curves class data write down the data answer questions on bottom of the sheet results of lab solubility of  $\text{KClO}_3$  unsaturated saturation and solubility curves in this experiment we expressed concentration in units of g solute g  $\text{H}_2\text{O}$  putting concentrations in g solute g solution would be tedious and impractical see fig 2

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web use the solubility curve to answer these questions determine how much potassium nitrate will dissolve in 20 g of water at 40 c 200 cm<sup>3</sup> of saturated lead ii nitrate solution was prepared at a temperature of 90 c

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web collect experimental data and create a solubility curve by the end of this lab students should be able to properly use an analytical balance to measure mass set up an experimental work station to measure the solubility of a salt in water as a function of the temperature generate a workbook using google sheets prior knowledge

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web may 29 2018 figure 2 the graph shows the solubility curve of potassium nitrate with corresponding temperatures as the temperature increases so does the solubility of the solution a y 0 2 1 61 b y 0 2 1 61 0 60 2 1 60 61 0 40 2 1 40 61 8 5 g 100 ml 4 6 g 100 ml

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web analysis construct a solubility curve using the mass of solute that will dissolve in 100 g of water at a specified temperature in this lab 5 ml of water was used assume 1 0 g ml of water take the collected data make it proportional to get the solubility of the kno3 graph temperature vs solubility using a best fit curve

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