

# CONCENTRATION OF SOLUTIONS



The concentration of a solution is usually measured in moles per cubic decimetre (mol/dm<sup>3</sup>). This is a measure of the number of moles in one cubic decimetre.

The volume must be in dm<sup>3</sup> (there are 1000 cm<sup>3</sup> in 1 dm<sup>3</sup>).  $\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$

$$\text{concentration (mol/dm}^3) = \frac{\text{moles}}{\text{volume (dm}^3)}$$



1) Calculate the concentration of the following solutions in mol/dm<sup>3</sup>.

a) 0.10 moles of NaCl in 200 cm<sup>3</sup>

b) 0.20 moles of H<sub>2</sub>SO<sub>4</sub> in 100 cm<sup>3</sup>

c) 0.500 moles of NaOH in 25 cm<sup>3</sup>

2) Calculate the number of moles in the following solutions.

a) 100 cm<sup>3</sup> of 0.20 mol/dm<sup>3</sup> HNO<sub>3</sub>

b) 25 cm<sup>3</sup> of 1.50 mol/dm<sup>3</sup> KOH

c) 50 cm<sup>3</sup> of 0.10 mol/dm<sup>3</sup> H<sub>2</sub>SO<sub>4</sub>

Concentration can also be measured in grams per cubic decimetre (g/dm<sup>3</sup>). This is a measure of the number of grams in one cubic decimetre. (Remember that mass = M = moles)

1 dm<sup>3</sup>

2 moles of H<sub>2</sub>SO<sub>4</sub>  
196 g. of H<sub>2</sub>SO<sub>4</sub>

Concentration = 2 mol/dm<sup>3</sup>

196 g. of H<sub>2</sub>SO<sub>4</sub> = 196

Concentration = 2 × 196 = 392 g/dm<sup>3</sup>

A simple conversion is:

$$\text{conc (g/dm}^3) = \text{conc (mol/dm}^3) \times M$$

3) Calculate the concentration of the following solutions in g/dm<sup>3</sup>.

a) 0.100 mol/dm<sup>3</sup> NaOH

b) 0.250 mol/dm<sup>3</sup> CH<sub>3</sub>COOH

c) 1.50 mol/dm<sup>3</sup> HNO<sub>3</sub>

4) 0.20 moles of NaCl is dissolved in 250 cm<sup>3</sup> of water.

a) Calculate the concentration in mol/dm<sup>3</sup>

b) Calculate the concentration in g/dm<sup>3</sup>

5) 5.0 g of KNO<sub>3</sub> is dissolved in 100 cm<sup>3</sup> of water.

a) Calculate the concentration in g/dm<sup>3</sup>

b) Calculate the concentration in mol/dm<sup>3</sup>

# Chemistry Solutions Study Guide Answers

**Susan McMurry, John McMurry**

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