

Chemistry Spring final exam review questions with answers below

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. A 1.0 M solution of hydrochloric acid conducts electricity better than a 1.0 M solution of acetic acid. Which of the following is the best explanation of this?

- HCl is a strong acid while $\text{HC}_2\text{H}_3\text{O}_2$ is weak.
- The solutions are not the same pH.
- Acetic acid has a higher molecular weight.
- Acetic acid has more ions in one formula unit.

2. The standard molar heat of vaporization for water is 40.79 kJ/mol. How much energy would be required to vaporize 94.0 g of water that started at 100°C and 1 atm pressure?

- 0.128 kJ
- 7.81 kJ
- 213 kJ
- 3800 kJ

3. What are the acids in the following equilibrium reaction?

$$\text{CN}^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCN}(\text{aq}) + \text{OH}^-(\text{aq})$$

- CN^- , OH^-
- H_2O , OH^-
- CN^- , H_2O
- H_2O , HCN

4. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change inside the container?

- The pressure will be five times higher.
- The pressure will be four times higher.
- The pressure will decrease by one fifth.
- The pressure will not change.

5. A sample of gas occupies 17 mL at -11.2°C . What volume does the sample occupy at 70°C ?

- 8.0mL
- 27 mL
- 36mL
- 10.6 mL

6. If the concentration for a solution is 4% by mass and the mass of the solution is 200 g, what is the mass of solute in solution?

- 8.0 g
- 50 g
- 800 g
- 80 g

7. The volume of 4.00M HCl that can be made from 3.19 mL of 6.80M HCl is_____.

- 281 mL
- 8.52 mL
- 5.42 mL
- 1.87 mL

8. What mass of sucrose (342 g/mol, $\text{C}_12\text{H}_{22}\text{O}_11$) is needed to make 400.0 mL of a 0.300M solution?

- 17.1 g
- 34.2 g
- 41.0 g
- 28.5 g

9. Sodium reacts with water according to the equation below. If 200. grams of sodium react, what volume of hydrogen gas is produced at 298K and 1.10 atm?

$$2 \text{Na}(\text{s}) + 2 \text{H}_2\text{O}(\text{l}) \rightarrow 2 \text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$$

- 96.8 L
- 48.4 L
- 193.6 L
- 2046 L

10. Which of the changes listed below would shift the following reaction to the left?

$$4 \text{HCl}(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{ClP}(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$$

- addition of Cl_2
- removal of Cl_2

Chemistry Spring Final Exam Review With Answers

Glenn Theodore Seaborg

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