

Name \_\_\_\_\_

# Matter and Phase Changes

## The Kinetic-Molecular Theory

The Kinetic-Molecular Theory (KMT) explains the behavior of gases based on the idea that gas particles are in constant motion. The key assumptions are:

1. Gases consist of large numbers of tiny particles that are far apart relative to their size.
2. Collisions between gas particles and between particles and container walls are elastic collisions.
3. Gas particles are in continuous, rapid, random motion.
4. There are no forces of attraction or repulsion between gas particles.
5. The average kinetic energy of gas particles depends on the temperature of the gas.

## Dalton's Law of Partial Pressure

Dalton's Law states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the individual gases. This law is also explained by the KMT: each gas in the mixture exerts its own pressure independently, and the total pressure is simply the sum of these individual pressures. Mathematically:  $P_{\text{total}} = P_1 + P_2 + P_3 + \dots + P_n$ , where  $P_{\text{total}}$  are the partial pressures of each gas in the mixture.

## Graham's Law of Diffusion Rates

Graham's Law states that the rate of diffusion of a gas is inversely proportional to the square root of its molar mass. According to Graham's law, lighter gases effuse faster than heavier gases. This can be explained by the KMT: lighter molecules have higher average speeds and therefore collide with the hole more frequently, leading to a faster rate of effusion.

This can be expressed as:

$$\frac{\text{Rate}_1}{\text{Rate}_2} = \sqrt{\frac{M_2}{M_1}}$$

where Rate<sub>1</sub> and Rate<sub>2</sub> are the diffusion rates of gases 1 and 2, and M<sub>1</sub> and M<sub>2</sub> are their molar masses.



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**Joacim Rocklöv**

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**Prentice Hall Chemistry**, 2000 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999

**Resources for Teaching Middle School Science** Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using

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