

Name _____

Matter and Phase Changes

The Kinetic-Molecular Theory

The Kinetic-Molecular Theory (KMT) explains the behavior of gases based on the idea that gas particles are in constant motion. The key assumptions are:

1. Gases consist of large numbers of tiny particles that are far apart relative to their size.
2. Collisions between gas particles and between particles and container walls are elastic collisions.
3. Gas particles are in continuous, rapid, random motion.
4. There are no forces of attraction or repulsion between gas particles.
5. The average kinetic energy of gas particles depends on the temperature of the gas.

Dalton's Law of Partial Pressure

Dalton's Law states that the total pressure of a mixture of gases is equal to the sum of the partial pressures of the individual gases. This law is also explained by the KMT: each gas in a mixture exerts its own pressure independently, and the total pressure is simply the sum of these individual pressures. Mathematically: $P_{\text{total}} = P_1 + P_2 + P_3 + \dots + P_n$ where P_{total} are the partial pressures of each gas in the mixture.

Graham's Law of Diffusion Rates

Graham's Law states that the rate of diffusion of a gas is inversely proportional to the square root of its molar mass. According to Graham's law, lighter gases effuse faster than heavier gases. This can be explained by the KMT: lighter molecules have higher average speeds and therefore collide with the hole more frequently, leading to a faster rate of effusion.

This can be expressed as:

$$\frac{\text{Rate}_1}{\text{Rate}_2} = \sqrt{\frac{M_2}{M_1}}$$

where Rate_1 and Rate_2 are the diffusion rates of gases 1 and 2, and M_1 and M_2 are their molar masses.



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Joacim Rocklöv



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Table of Contents Chemistry Study Guide Answers Phase And Changes

1. Understanding the eBook Chemistry Study Guide Answers Phase And Changes
 - The Rise of Digital Reading Chemistry Study Guide Answers Phase And Changes
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Answers Phase And Changes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Answers Phase And Changes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Answers Phase And Changes
 - Personalized Recommendations
 - Chemistry Study Guide Answers Phase And Changes User Reviews and Ratings
 - Chemistry Study Guide Answers Phase And Changes and Bestseller Lists
5. Accessing Chemistry Study Guide Answers Phase And Changes Free and Paid eBooks
 - Chemistry Study Guide Answers Phase And Changes Public Domain eBooks
 - Chemistry Study Guide Answers Phase And Changes eBook Subscription Services
 - Chemistry Study Guide Answers Phase And Changes Budget-Friendly Options
6. Navigating Chemistry Study Guide Answers Phase And Changes eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Study Guide Answers Phase And Changes Compatibility with Devices
- Chemistry Study Guide Answers Phase And Changes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Answers Phase And Changes
 - Highlighting and Note-Taking Chemistry Study Guide Answers Phase And Changes
 - Interactive Elements Chemistry Study Guide Answers Phase And Changes
- 8. Staying Engaged with Chemistry Study Guide Answers Phase And Changes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Answers Phase And Changes
- 9. Balancing eBooks and Physical Books Chemistry Study Guide Answers Phase And Changes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Answers Phase And Changes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide Answers Phase And Changes
 - Setting Reading Goals Chemistry Study Guide Answers Phase And Changes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide Answers Phase And Changes
 - Fact-Checking eBook Content of Chemistry Study Guide Answers Phase And Changes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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