



Reaction Rates and Equilibrium Study Guide

Topics Covered:

Factors affecting reaction rates

Collision theory of reaction rates

Reaction progress diagrams (activation energy and transition states)

Rate laws and the method of initial rates

Chemical equilibrium

Equilibrium expressions and constants and their meaning

Product-favored vs. Reactant-favored

Shifting equilibrium and Le Chatelier's Principle

Quantitative calculations of equilibrium systems using ICE tables

Practice Questions: [Click here for an answer key](#)

1. Describe how each of the following factors would affect the rate of "particle collisions" in your response.
 - a. Temperature decrease
 - b. Addition of iodide ions, I^- , as a catalyst
 - c. Increased reactant concentration
 - d. Crushing a solid into a powder (with larger surface area)
2. The reaction progress diagram shows the energy changes that take nitrogen dioxide gas.



Chemistry Study Guide Answers Reaction Rates

Konstantinos Papadopoulos

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