

TEACHER GUIDE AND ANSWERS

Study Guide - Chapter 10 – The Mole

Section 10.1 Measuring Matter

1. pair
2. 5
3. dozen
4. gross
5. 200
6. ream
7. 6,000,000,000
8. 0.5 mol
9. 6.02×10^{23}
10. four moles
11. $\frac{6.02 \times 10^{23} \text{ Cu atoms}}{1 \text{ mol Cu}}$
12. $\frac{1 \text{ mol CH}_4}{6.02 \times 10^{23} \text{ molecules CH}_4}$
13. $\frac{1 \text{ mol Xe}}{6.02 \times 10^{23} \text{ molecules Xe}}$
14. $\frac{6.02 \times 10^{23} \text{ molecules F}_2}{1 \text{ mol F}_2}$

Section 10.2 Mass and the Mole

1. false
2. true
3. false
4. true
5. true
6. true
7. false
8. true
9. b
10. d, a
11. a
12. b, c
13. d
14. c

Section 10.3 Moles of Compounds

1. carbon, 1 atom; hydrogen, 4 atoms
2. carbon, 1 atom; hydrogen, 1 atom; chlorine, 3 atoms
3. 1 mol C; 4 mol H
4. 1 mol C; 1 mol H; 3 mol Cl
5. 6.02×10^{23}
6. 1.81×10^{24}
7. 16.05 g/mol
8. 1.20×10^{24}

Section 10.4 Empirical and Molecular Formulas

1. The percent composition of a compound is the percent by mass of each of the elements in a compound.
2. Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100.
3. c
4. c
5. b
6. b
7. c
8. c
9. d
10. $7.89 \text{ g K} \times 1 \text{ mol K} / 39.10 \text{ g K} = 0.202 \text{ mol K}$
 $2.42 \text{ g C} \times 1 \text{ mol C} / 12.01 \text{ g C} = 0.201 \text{ mol C}$
 $9.69 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 0.606 \text{ mol O}$
 $0.202 \text{ mol K} / 0.201 \text{ mol C} = 1.00 \text{ mol K} / 1.00 \text{ mol C} = 1 \text{ mol K} / 1 \text{ mol C}$
 $0.201 \text{ mol C} / 0.201 \text{ mol C} = 1.00 \text{ mol C} / 1.00 \text{ mol C} = 1 \text{ mol C} / 1 \text{ mol C}$
 $0.606 \text{ mol O} / 0.201 \text{ mol C} = 3.01 \text{ mol O} / 1.00 \text{ mol C} = 3 \text{ mol O} / 1 \text{ mol C}$
1 mol K : 1 mol C : 3 mol O
empirical formula: KCO_3
(39.10 g + 12.01 g + 48.00 g) = 99.11 g

Chapter 10 The Mole Study Guide Answers

**William K. Purves, Edward
Dzialowski, Lindsay Goodloe, Betty
McGuire, Nancy Guild, Paula Mabee**

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Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry

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