

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

11. A car dealer has 100 bags of apples, 20 crates of peaches, and 10 tubs of peaches left over. Write a balanced equation that fits this situation. How many apples are needed to make one peach?
12. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $2\text{KClO}_3(s)$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
13. Using the equation from problem 12, how many moles of oxygen are produced from 1.0 moles of potassium chlorate?
14. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 2.0×10^{24} molecules of oxygen? How many moles of water are produced from 22.0×10^{24} molecules of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

15. Calculate the number of moles of hydrogen chloride produced from 30 moles of hydrochloric acid.



16. Calculate the number of moles of chlorine needed to form 1.0 moles of iron(III) chloride.



17. Calculate the number of grams of nitrogen dioxide that are produced from 4.0 moles of nitric oxide.



18. Calculate the mass, in grams, produced from the decomposition of 75.0 g of potassium chlorate.



19. Calculate the mass, in grams, needed to react with chlorine to produce 164 g of silver chloride. (Refer to Table 1 for the molar mass of each element.)

20. How many liters of carbon monoxide at 500°C are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



21. Calculate the number of moles of oxygen gas needed to produce 4.00 liters of nitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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James Joseph Gallagher

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