

## Chapter 18

### Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{\text{eq}} = \frac{[\text{Pb}^{2+}][\text{F}^-]^2}{[\text{PbF}_2]}$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}$ ;  $[\text{HCl}] = 1.0 \times 10^{-2}$

$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{\text{eq}} = \frac{(1.0 \times 10^{-2})^2}{(1.0 \times 10^{-2})(1.0 \times 10^{-2})} = 0.0001$$



$[\text{N}_2] = 4.4 \times 10^{-2}$ ;  $[\text{H}_2] = 1.2 \times 10^{-1}$ ;  $[\text{NH}_3] = 3.4 \times 10^{-3}$

$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{\text{eq}} = \frac{(3.4 \times 10^{-3})^2}{(4.4 \times 10^{-2})(1.2 \times 10^{-1})^3} = 0.152$$

3. For the reaction  $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ ,  $K_{\text{eq}} = 0.2$ . At a particular time, the following concentrations are measured:  $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$ ,  $[\text{NO}_2] = 0.2 \text{ M}$ . Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$

$[\text{NO}_2] = 0.2 \text{ M}$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$Q = \frac{(0.2)^2}{(2.0)} = 0.02$$

$K_{\text{eq}} = 0.2$

$Q < K_{\text{eq}}$ , so the reaction will proceed to the right (forward).

# Chapter 18 Study Guide Acids And Bases

**Neil E. Schore**

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