

**CHAPTER 15** STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

After reading (1) \_\_\_\_\_ solution \_\_\_\_\_ of solute is dissolved in solvent.

(2) \_\_\_\_\_ solute \_\_\_\_\_ is the (3) \_\_\_\_\_ solvent \_\_\_\_\_ and solute is the

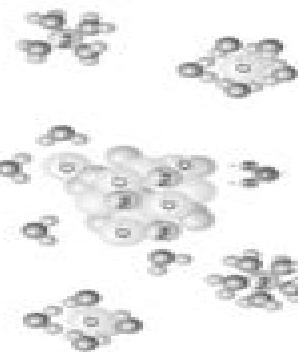
(4) \_\_\_\_\_ solution. A solution that does not dissolve is  
 (5) \_\_\_\_\_ insoluble. (6) \_\_\_\_\_ liquid \_\_\_\_\_ solution is the most common  
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the  
 two liquids are (7) \_\_\_\_\_ miscible. However, if one liquid is insoluble in another  
 the liquids are (8) \_\_\_\_\_ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.



1. Indicate in solute is when the solute is in a solution. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 15

Name \_\_\_\_\_ Date \_\_\_\_\_

**CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY****Section 15.1** *Answers*

16. An sodium chloride dissolved in water. What happens is the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecule around the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved, stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with type of solute.

True 14. The overall energy change that occurs when a solute dissolves is heat of solution.

False 15. Solubility is a property of the solute, amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False 16. Solvent molecules in liquid as the solute has a low heat of solution.

True 17. In a saturated solution, solute and solvents are in equilibrium.

True 18. Additional solute can be dissolved in an unsaturated solution.

True 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Chapter 4 Study Guide For Content Mastery Answer Key

**Kenneth H. Hoover**

## **Chapter 4 Study Guide For Content Mastery Answer Key:**

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## **Table of Contents Chapter 4 Study Guide For Content Mastery Answer Key**

1. Understanding the eBook Chapter 4 Study Guide For Content Mastery Answer Key
  - The Rise of Digital Reading Chapter 4 Study Guide For Content Mastery Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Study Guide For Content Mastery Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 4 Study Guide For Content Mastery Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Study Guide For Content Mastery Answer Key
  - Personalized Recommendations
  - Chapter 4 Study Guide For Content Mastery Answer Key User Reviews and Ratings
  - Chapter 4 Study Guide For Content Mastery Answer Key and Bestseller Lists
5. Accessing Chapter 4 Study Guide For Content Mastery Answer Key Free and Paid eBooks
  - Chapter 4 Study Guide For Content Mastery Answer Key Public Domain eBooks
  - Chapter 4 Study Guide For Content Mastery Answer Key eBook Subscription Services
  - Chapter 4 Study Guide For Content Mastery Answer Key Budget-Friendly Options
6. Navigating Chapter 4 Study Guide For Content Mastery Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Chapter 4 Study Guide For Content Mastery Answer Key Compatibility with Devices
- Chapter 4 Study Guide For Content Mastery Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Chapter 4 Study Guide For Content Mastery Answer Key
- Highlighting and Note-Taking Chapter 4 Study Guide For Content Mastery Answer Key
- Interactive Elements Chapter 4 Study Guide For Content Mastery Answer Key

8. Staying Engaged with Chapter 4 Study Guide For Content Mastery Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chapter 4 Study Guide For Content Mastery Answer Key

9. Balancing eBooks and Physical Books Chapter 4 Study Guide For Content Mastery Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Chapter 4 Study Guide For Content Mastery Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Chapter 4 Study Guide For Content Mastery Answer Key

- Setting Reading Goals Chapter 4 Study Guide For Content Mastery Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 4 Study Guide For Content Mastery Answer Key

- Fact-Checking eBook Content of Chapter 4 Study Guide For Content Mastery Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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